

Reaction Quotient

How do you predict which direction a reaction will proceed to reach equilibrium?

Why?

When a reaction reaches equilibrium there must be some non-negligible amount of every species in the reaction, otherwise the reaction cannot react in both directions. Knowing this, it is very easy to predict which direction a reaction will go to reach equilibrium when one of the components of the reaction has an initial concentration of zero. Many of the problems you have worked with thus far have some component at zero concentration, but real life does not work that way. Most of the time, the reaction in question has some measurable amount of every species. Deciding which way the reaction will go to reach equilibrium then becomes more challenging.

Model 1 – A Theoretical Equilibrium

Trial 1	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.000 M	1.000 M	1.000 M
Change			
Equilibrium	1.464 M	1.464 M	0.536 M

Trial 3	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.000 M	0.500 M	1.500 M
Change			
Equilibrium	1.864 M	1.364 M	0.636 M

Trial 5	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.400 M	1.200 M	0.400 M
Change			
Equilibrium	1.388 M	1.188 M	0.412 M

Trial 2	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	2.000 M	0.500 M	0.500 M
Change			
Equilibrium	2.150 M	0.650 M	0.350 M

Trial 4	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.600 M	1.000 M	0.400 M
Change			
Equilibrium	1.600 M	1.000 M	0.400 M

Trial 6	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	0.750 M	2.000 M	0.250 M
Change			
Equilibrium	0.675 M	1.925 M	0.325 M

- Examine Model 1.
 - Write the theoretical chemical reaction that is used in the trials of Model 1.
 - If 0.50 M of reactant A reacts, predict the change in concentration of B and C.
- What variables were changed in the different trials shown in Model 1?

Pogil Activities For High School Chemistry Equilibrium

Laying the Foundation

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

Pogil Activities For High School Chemistry Equilibrium:

Questioning for Formative Feedback Jackie A. Walsh, 2022-05-20 When used effectively quality questions and student dialogue result in self regulated learners and formative feedback that reveals progress toward learning goals Learning knows no boundaries The potential for learning exists whenever and wherever we interact with our environment So how can we infuse school learning with the authenticity and excitement associated with real life experiences In Questioning for Formative Feedback Jackie Acree Walsh explores the relationship between questioning and feedback in K 12 classrooms and how dialogue serves as the bridge connecting the two Quality questioning productive dialogue and authentic use of feedback are a powerful trifecta for addressing the needs of a new generation of learners In fact the skillful use of these three processes can fuel and accelerate the academic social and emotional learning of all students In this book Walsh provides a manual of practice for educators who want to engage students as partners in these processes To that end she offers the following features to help create a classroom in which everyone learns through intentional practice Blueprints for coherent models of key processes and products Tools and strategies to help you achieve identified outcomes Protocols with step by step directions to complete an activity Classroom artifacts of authentic classroom use including links to 21 original videos produced exclusively for this book Working together questioning dialogue and feedback can transform learning for all This book supports you in embracing and bringing that vision to fruition *POGIL Activities for High School Chemistry* High School POGIL Initiative, 2012

High School Chemistry Vishal Mody, 2016-05-20 This chemistry booklet was created to help students specifically with the topic of Chemical Equilibrium in chemistry This booklet has been made extremely concise yet explains the concepts in detail at the same time Remember that this booklet is not designed to be your main study source but rather as an adjunct to your school teacher s notes There are also lots of practice questions with detailed solutions at the end to solidify the concepts you have learned **Chemical Equilibrium In a Nutshell** Brian Coleman, 2018-06-25

Understanding the math and minutiae of chemical equilibrium can be a tall task for anyone so why not enlist the help of a scientific squirrel to guide you on your journey Join Dr Wash as we dabble in equilibrium constants and other tools needed to predict chemical processes This book focuses on introductory concepts at the high school and early university level focusing on identifying equilibrium calculating K and Q discussing Le Chatelier s principle and tying equilibrium with the field of Thermodynamics Full of step by step instructions and practice questions this book aims to simplify one of the more complex topics found within the field of chemistry **Chemical Equilibrium** William Guenther, 1975-12 The present work is

designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions

of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions. There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions. Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations. The coverage of solution equilibrium phenomena is not however exhaustively comprehensive. Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria. Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation. **Equilibrium Calculations** John T. Donoghue, 1971

Interactive Chemistry, 2001 *Principles of Chemical Equilibrium* Brian Craven, 1972 **Equilibrium** William Edward Marsden, 1996 **Principles of Chemical Equilibrium** Brian Raymond Craven, 1972 Principles of Chemical Equilibrium Kelso B. Morris, 1965 **Chemical Equilibrium** Robert Blanchard Fischer, Dennis G. Peters, 1970

Equilibrium and Energetics John Allen Berry-Richards, 1972 *Molecular Equilibrium* P. H. Carnell, R. N. Reusch, 1963

Instant Lessons in Chemistry Denise De Vreeze, Kath McMicking, 1998 This publication consists of reproducible worksheets usually two pages suitable for senior high school Chemistry. The worksheets aim to give students experience in applying concepts interpreting and presenting data and building a core of chemistry knowledge. **Equilibrium III**, 1984

Equilibrium Laying the Foundation, 2010-05-15 *CHEMICAL EQUILIBRIUM* NARAYAN CHANGDER, 2024-04-01 Note: Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u@gmail.com. You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz>. I will send you a PDF version of this workbook. This book has been designed for candidates preparing for various competitive examinations. It contains many objective questions specifically designed for different exams. Answer keys are provided at the end of each page. It will undoubtedly serve as the best preparation material for aspirants. This book is an engaging quiz eBook for all and offers something for everyone. This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information. Use this invaluable book to test your subject matter expertise. Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment. Although the majority of students are accustomed to this MCQ format, many are not well versed in it. To achieve success in MCQ tests, quizzes, and trivia challenges, one requires test-taking techniques and skills in addition to subject knowledge. It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations. Whether you have studied the subject on your own, read for pleasure, or completed coursework, it will assess your knowledge and prepare you for competitive exams, quizzes, trivia, and more. **Equilibrium** Laying the Foundation, 2010-05-15 **Chemical Equilibrium and Solutions** James E. Banks, 1967

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