

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, S, Cl, O

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

O, Cl, S, Na, K

3. Which atom has a higher electronegativity: fluorine or chlorine?

Fluorine is more electronegative than chlorine because it is smaller.

4. Which two elements in the same family generally have similar properties?

There have been many cases where atoms in the same family have similar properties.

5. Arrange the following properties in order from left to right across the periodic table:

- a. atomic radius (increases to the right) **increases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest atomic radius?

Period 7, francium, cesium, barium, lanthanum, actinides, and the lanthanide series.

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest atomic radius?

Period 2, fluorine, neon, oxygen, nitrogen, carbon, boron, and beryllium.

8. Write the atomic radius for the following elements:

Li	152
Be	112
B	85
C	77
N	75
O	73
F	71
Ne	71

Na	186
Mg	145
Al	118
Si	111
P	106
S	104
Cl	99
Ar	98

9. Write the atomic radius for the following elements:

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