

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? _____ (oxygen)
which cation of calcium always forms with a charge of 2+? _____ (calcium)
2. How electrons in the halogen group has _____ valence electrons, and needs to _____ 1 more to have 8. This is because of the _____ rule.
3. How _____ bond in the alkaline elements oppositely charged ions.
How is ionic bond is _____
4. On a molecular level, atoms _____ molecules, formed by an ionic bond, atoms _____ molecules.
5. Elements group 16 has 6 valence electrons. It is usually found as _____, and atoms group 17 usually found as _____.
How can group 17 be distinguished, in terms, other elements, or molecules formed by ionic bonding? _____
6. How atoms can have up to _____ valence electrons. This usually happens in the periodic table that all have 8 valence electrons, or called _____
7. Atoms can gain, lose, or share _____ by forming ionic bonds.

B. Formulas for Ionic Compounds - Fill in the chart below

Write the correct chemical formula for each of the following	Magnesium	Fluorine	Phosphorus	Chlorine
1. Lithium				
2. Sodium				
3. Potassium				
4. Calcium				
5. Barium				
6. Strontium				
7. Radium				
8. Beryllium				
9. Magnesium				
10. Zinc				
11. Cadmium				
12. Mercury				
13. Silver				
14. Gold				
15. Copper				
16. Nickel				
17. Cobalt				
18. Iron				
19. Manganese				
20. Chromium				
21. Vanadium				
22. Titanium				
23. Zirconium				
24. Niobium				
25. Molybdenum				
26. Technetium				
27. Ruthenium				
28. Rhodium				
29. Palladium				
30. Silver				
31. Cadmium				
32. Indium				
33. Tin				
34. Lead				
35. Bismuth				
36. Polonium				
37. Astatine				
38. Tellurium				
39. Selenium				
40. Sulfur				
41. Tellurium				
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93. Selenium				
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96. Selenium				
97. Sulfur				
98. Tellurium				
99. Selenium				
100. Sulfur				

3. Write the chemical formula for each of the following ions:
Cation: _____

Write your own ionic compounds below the line, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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