

Chapter 5 • Periodicity Activity

ATOMIC RADII

Directions: Using the given data for the sizes of atoms (atomic radii) below,

1. Create a line graph (use a ruler) by graphing *Period 2's* data on the graph below.
2. In a different color, create a line graph *Period 3's* data on the same graph.
3. Label each axis, and include a key/legend.
4. Answer all questions.

	1	2	13	14	15	16	17	18
1	H 1	Atomic Radii						He 4
2	Li 123	Be 89	B 88	C 77	N 70	O 66	F 64	Ne —
3	Na 157	Mg 136	Al 125	Si 117	P 110	S 104	Cl 99	Ar —
4	K 203	Ca 174	Ga 125	Ge 122	As 121	Se 117	Br 114	Kr —
5	Rb 216	Sr 192	In 150	Sn 140	Sb 141	Te 137	I 133	Xe —
6	Cs 235	Ba 198	Tl 155	Pb 154	Bi 152	Po 153	At —	Rn —

Note: Mark the vertical axis from 0 – 160 (by 20's)

Trend in Atomic Size Across the Periodic Table

Li	Be	B	C	N	O	F
Na	Mg	Al	Si	P	S	Cl

Horizontal Trend Questions

1. What is the trend in atomic size as you go ~~across~~ the periodic table? _____
2. This is true because there are more _____ (protons / electrons / layers of electrons) pulling the electron cloud in toward the nucleus.

Vertical Trend Questions

4. This is true because there are more _____ (protons / electrons / layers of electrons) making the electron cloud larger.

Periodicity Activity Answers

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Table of Contents Periodicity Activity Answers

1. Understanding the eBook Periodicity Activity Answers
 - The Rise of Digital Reading Periodicity Activity Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodicity Activity Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodicity Activity Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodicity Activity Answers
 - Personalized Recommendations
 - Periodicity Activity Answers User Reviews and Ratings
 - Periodicity Activity Answers and Bestseller Lists
5. Accessing Periodicity Activity Answers Free and Paid eBooks
 - Periodicity Activity Answers Public Domain eBooks

- Periodicity Activity Answers eBook Subscription Services
- Periodicity Activity Answers Budget-Friendly Options
- 6. Navigating Periodicity Activity Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodicity Activity Answers Compatibility with Devices
 - Periodicity Activity Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodicity Activity Answers
 - Highlighting and Note-Taking Periodicity Activity Answers
 - Interactive Elements Periodicity Activity Answers
- 8. Staying Engaged with Periodicity Activity Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodicity Activity Answers
- 9. Balancing eBooks and Physical Books Periodicity Activity Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodicity Activity Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodicity Activity Answers
 - Setting Reading Goals Periodicity Activity Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodicity Activity Answers
 - Fact-Checking eBook Content of Periodicity Activity Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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