

Periodic Trends Worksheet

Using a few minutes before today's activity, complete the following questions. Please post your responses on the board when you are done.

Additional help on the periodic trends can be found on this video: <https://www.youtube.com/watch?v=0D8Wp-2GQp0>

1. What does electronegativity measure and how does it change across the periodic table? (electronegativity is largest) in which direction, left to right, top to bottom, and bottom to top?
2. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
3. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. Fluorine is Li
 - b. It is not
 - c. Cl is not
4. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. It is not
 - b. Calcium is
 - c. Li is not
5. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. Cl is not
 - b. Fluorine is
 - c. It is not
6. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. Fluorine is
 - b. It is not
 - c. It is not
7. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. It is not
 - b. It is not
 - c. It is not
8. Atomic radius measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. It is not
 - b. It is not
 - c. It is not
9. Atomic radius measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. It is not
 - b. It is not
 - c. It is not
10. Atomic radius measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. It is not
 - b. It is not
 - c. It is not

Periodic Trends Questions And Answers

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