

Periodic Trends

Can the properties of an element be predicted using a periodic table?

Why?

The periodic table is often considered to be the "best friend" of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

3 1. Consider the data in Model 1 on the following page.

a. Each element has three numbers listed under it. Which value represents the atomic radius?

The first number below the diagram

b. What are the units for the atomic radius?

pm picometers 10^{-12} m

c. Write a complete sentence to convey your understanding of atomic radius. *Note:* You may not use the word "radius" in your definition.

Distance from center of atom (nucleus) to outer electrons

2 2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.

Increases $H < Li < Na < K$
 $C < Si < Ge$

$F < Cl < Br$



2 3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2. *Hint:* You should discuss either a change in distance between the nucleus and outer shell of electrons or a change in the number of protons in the nucleus.

As the number of electrons increases, new energy levels need to be added further from the nucleus, increasing the radius.

2 4. In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods.

Decreases $Li > Be > B > C > N > O > F > Ne$
 $Na > Mg > Al > Si > P > S > Cl > Ar$



2 5. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.

Highest energy level is constant, but the number of protons in the nucleus increases, increasing the charge of the nucleus and the attractive force on the electrons, pulling the highest energy level closer to the nucleus.

Periodic Trends Packet Answer Key

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Table of Contents Periodic Trends Packet Answer Key

1. Understanding the eBook Periodic Trends Packet Answer Key
 - The Rise of Digital Reading Periodic Trends Packet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Packet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Packet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Packet Answer Key
 - Personalized Recommendations
 - Periodic Trends Packet Answer Key User Reviews and Ratings
 - Periodic Trends Packet Answer Key and Bestseller Lists
5. Accessing Periodic Trends Packet Answer Key Free and Paid eBooks

- Periodic Trends Packet Answer Key Public Domain eBooks
 - Periodic Trends Packet Answer Key eBook Subscription Services
 - Periodic Trends Packet Answer Key Budget-Friendly Options
6. Navigating Periodic Trends Packet Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Trends Packet Answer Key Compatibility with Devices
 - Periodic Trends Packet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Trends Packet Answer Key
 - Highlighting and Note-Taking Periodic Trends Packet Answer Key
 - Interactive Elements Periodic Trends Packet Answer Key
8. Staying Engaged with Periodic Trends Packet Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Packet Answer Key
9. Balancing eBooks and Physical Books Periodic Trends Packet Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Packet Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Packet Answer Key
- Setting Reading Goals Periodic Trends Packet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Packet Answer Key
- Fact-Checking eBook Content of Periodic Trends Packet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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