

## The Periodic Law Activity

Name/Number: \_\_\_\_\_

**Introduction** The present organization of the elements is a product of the first periodic table published by Dmitri Mendeleev in 1869. The amazing accuracy of his predictions has been very important to chemists in this century. However, the basis of his arrangement was the atomic mass of the elements. This approach proved incorrect as it would have placed some elements in a family with dissimilar properties. Henry Moseley rearranged the table on the basis of the atomic numbers of the elements. In accordance with Moseley's revision, the **periodic law** states: *the properties of the elements are periodic functions of their atomic numbers*. Each of the 115 known elements has a unique set of characteristic properties. These range from solid to gas, lustrous to dull, low to high melting points, various colors, and so on. The elements are arranged within the periodic table into groups or families (vertical columns) and periods (horizontal rows). This arrangement reflects the periodic or repeating nature of the properties of the elements.

**Learning Target** Students will describe the following trends: atomic radii, ionization energy and electronegativity.

**What the Numbers in the boxes mean**

Atomic #	Code letter	Symbol
Ionization Energy (kJ/mol)		Phase
Atomic radius (pm)		Electronegativity

### Procedure

- Each block on the table represents a different element from the first 4 periods of Groups 1, 2 and 13-18.
- Cut out blocks A - Z.
- The following sets of elements are in families: ZRD, PSIF, JXBE, LHT, QKA, WOV, GUN, and YMC.
- Use the following clues and arrange the elements in their proper order on the blank table provided.
 

J has an atomic number three times that of T.  
 The atomic mass of L is 40.08 amu.  
 U has a total of six electrons.  
 S is an alkali metal.  
 E is a noble gas.  
 W is a liquid.  
 Z has the smallest atomic mass in its group.  
 B has ten protons.  
 I has the largest atomic number in its group.  
 O has an atomic number larger than V.  
 D has the largest atomic mass of its group.  
 X has an atomic number one higher than F.

L is an alkaline earth metal.  
 S forms a compound with O that is used as a seasoning on food.  
 Y is a metalloid.  
 O is a halogen.  
 The atomic mass of T is more than that of H.  
 Q has an atomic mass 2 times that of A.  
 M has an atomic number one less than that of A.  
 The electrons of N are distributed over three energy levels.  
 R has 13 protons in its nucleus.  
 A is essential for all animals to breathe.
- When you have placed these 26 elements in their correct position, glue them in place.
- Check with your instructor for the correct answer.

**Questions** Answer questions 2 and 3 in complete sentences.

- In each block, fill in the element's atomic number in the upper left-hand corner and the element's correct symbol in the upper right hand corner. Attach this table to the FRONT of your lab when you turn it in.
- Define each of the following terms. Then examine your completed table. What *general observations* can be made of trends within *periods and groups* for the following properties? In your answer, include whether the following values tend to increase or decrease across the period and then whether it increases or decreases down a group. Please look at going left to right across the period and from top to bottom in the group.
  - Ionization Energy

# Periodic Law Lab Answers

**LP Steffe**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'LP Steffe' text.

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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures *The Science Teachers Bulletin* ,1989 **The OHA Law Reporter** ,1985 *The Cambridge Handbook of Law and Entrepreneurship in the United States* D. Gordon Smith, Brian Broughman, Christine Hurt, 2022-04-14 Law plays a key role in determining the level of entrepreneurial action in society Legal rules seek to define property rights facilitate private ordering and impose liability for legal wrongs thereby attempting to establish conditions under which individuals may act These rules also channel the development of technology regulate information flows and determine parameters of competition Depending on their structure and implementation legal rules can also discourage individuals from acting It is thus crucial to determine which legal rules and institutions best enable entrepreneurs whose core function is to challenge incumbency This volume assembles legal experts from diverse fields to examine the role of law in facilitating or impeding entrepreneurial action Contributors explore issues arising in current policy debates including the incentive effect of legal rules on startup activity the role of law in promoting or foreclosing market entry and the effect of entrepreneurial action on legal doctrine *Interactive School Science 10* , The Software Encyclopedia 2000 Bowker Editorial Staff, 2000-05 Chemistry William Strubhar Haldeman, 1920 **Science** John Michels (Journalist), 1918 Vols for 1911 13 contain the Proceedings of the Helminothological Society of Washington ISSN 0018 0120 1st 15th meeting *Jail & Prisoner Law Bulletin* ,1994 **Hot Matter from High-Power Lasers** Peter Mulser, 2020-08-01 This book offers an introduction to the booming field of high power laser matter interaction It covers the heating of matter to super high temperatures and pressures novel schemes of fast particle acceleration matter far from thermal equilibrium stimulated radiation scattering relativistic optics strong field QED as well as relevant applications such as extreme states of matter controlled fusion and novel radiation sources All models and methods considered are introduced as they arise and illustrated by relevant examples Each chapter contains a selection of problems to test the reader's understanding to apply the models under discussion to relevant situations and to discover their limits of validity The carefully chosen illustrations greatly facilitate the visualization of physical processes as well as presenting detailed numerical results A list of useful formulas and tables are provided as a guide to quantifying results from experiments and numerical simulations Each chapter ends with a description of the state of the art and the current research frontiers **Energy Research Abstracts** ,1978-12 *Human Engineering Bibliography* ,1958 *SEC Docket* United States. Securities and Exchange Commission, 2008 Scientific and Technical Aerospace Reports ,1994-06 *Scientifica* David Sang, Peter Ellis, Derek McMonagle, 2004 Bring your science lessons to life with

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