#### **GRAPHING PERIODIC TRENDS**

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.

Define the following terms(starting on page 170):

Atomic radius-

First ionization energy-

Electronegativity-

#### Graph 1 - Atomic Radius as a function of Atomic Number

- A. Create a graph of the atomic radius as a function of atomic number. Plot atomic number on the X axis and atomic radius on the Y axis. Remember to label the axes!
- B. Use a colored pen, pencil or highlighter to trace over the element's period (horizontal row on the periodic table). For example: use GREEN to trace for all of the elements in row 1, then use YELLOW to trace for all of the elements in row 2, then use ORANGE to trace for all the elements in row 3, then use BLUE to trace for all the elements in row 4.

Symbol	Atomic Number	Atomic Radius(pm)					
H	1	37					
He	2	50					
Li	3	123					
Be	4	89					
В	5	80					
C	6	77					
N	7	70					
0	8	66					
F.	9	64					
Ne	10	67					
Na	1.1	157					
Mg	12	136					
AI	13	125					
Si	1.4	1.17					
P	1.5	110					
S	16	104					
CI	17	99					
Air	18	98					
K	19	203					
Ca	20	174					

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- Describe the trend in atomic radius as the atomic number increases across a period.
- 2. Describe the trend in atomic radius as the atomic number increases down a group.

# **Periodic Trends Answers Extension Questions**

Tina Overton, Jonathan Rourke, Fraser A. Armstrong

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