

# The Periodic Table of Elements

In 1869 Dimitri Mendeleev organized the elements in a table based on the properties of the elements.

Groups

↓ ↓  
(Columns)

The elements in each group have the same number of valence electrons.

|                                          |                                             |                   |                                    |                    |                                    |                   |                                    |                    |                                    |                    |                                    |                   |                                    |                   |                                    |                  |                                               |
|------------------------------------------|---------------------------------------------|-------------------|------------------------------------|--------------------|------------------------------------|-------------------|------------------------------------|--------------------|------------------------------------|--------------------|------------------------------------|-------------------|------------------------------------|-------------------|------------------------------------|------------------|-----------------------------------------------|
| 1 <sup>st</sup> Period → (1 shell) ☉     | Group 1-2<br>(1, 2 valence e <sup>-</sup> ) | 1 H<br>Hydrogen   | Group 2-2<br>(2 valence electrons) | 4 Be<br>Beryllium  | Group 3-2<br>(3 valence electrons) | 5 B<br>Boron      | Group 4-2<br>(4 valence electrons) | 6 C<br>Carbon      | Group 5-2<br>(5 valence electrons) | 7 N<br>Nitrogen    | Group 6-2<br>(6 valence electrons) | 8 O<br>Oxygen     | Group 7-2<br>(7 valence electrons) | 9 F<br>Fluorine   | Group 8-2<br>(8 valence electrons) | 2 He<br>Helium   | ↳ Except Helium with 2 valence e <sup>-</sup> |
| 2 <sup>nd</sup> Period → (2 shells) ☉☉   |                                             | 3 Li<br>Lithium   |                                    |                    |                                    |                   |                                    |                    |                                    |                    |                                    |                   |                                    |                   |                                    | 10 Ne<br>Neon    |                                               |
| 3 <sup>rd</sup> Period → (3 shells) ☉☉☉  |                                             | 11 Na<br>Sodium   |                                    | 12 Mg<br>Magnesium |                                    | 13 Al<br>Aluminum |                                    | 14 Si<br>Silicon   |                                    | 15 P<br>Phosphorus |                                    | 16 S<br>Sulfur    |                                    | 17 Cl<br>Chlorine |                                    | 18 Ar<br>Argon   |                                               |
| 4 <sup>th</sup> Period → (4 shells) ☉☉☉☉ |                                             | 19 K<br>Potassium |                                    | 20 Ca<br>Calcium   |                                    | 31 Ga<br>Gallium  |                                    | 32 Ge<br>Germanium |                                    | 33 As<br>Arsenic   |                                    | 34 Se<br>Selenium |                                    | 35 Br<br>Bromine  |                                    | 36 Kr<br>Krypton |                                               |

Periods →  
(rows)

The elements in each period have the same number of shells.

Today, 114 elements are arranged by atomic number in the periodic table. The atomic number is based on the number of protons in an element.

# Periodic Table Notes For Middle School

**Naomi Silver, Matthew Kaplan, Danielle  
LaVaque-Manty, Deborah Meizlish**

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