

Atoms are most stable when they achieve the number of electrons necessary to reach a **Noble Gas Configuration**.

□ The tendency for atoms to lose or gain electrons in order to reach this configuration is known as the Octet rule

		The Organic Table of Elements																									
Group		1A	2A																			3A	4A	5A	6A	7A	8A
Period																											
1		1 H																									2 He
2		3 Li	4 Be																			5 B	6 C	7 N	8 O	9 F	10 Ne
3		11 Na	12 Mg																			13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
		19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr								
		37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe								
		55 Cs	56 Ba	71 Lu	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn								
		87 Fr	88 Ra	103 Lr	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og								

Metals

Metalloids

Nonmetals

Group 8A is highlighted with a red circle and an arrow pointing to Helium (He).





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