

The Periodic Law

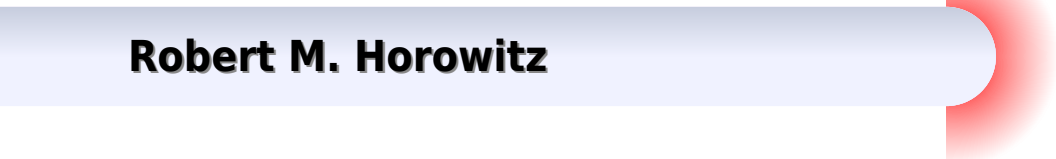
SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave better chemical properties of elements appeared in regular intervals and whereas elements were arranged according to increasing
 - (a) density
 - (b) atomic number
 - (c) atomic mass
 - (d) electronegativity
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of their atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a new period
 - (c) adding a new column
 - (d) adding a new group
5. _____ The main reason the properties of the noble gases is that they are
 - (a) inert
 - (b) unreactive
 - (c) large
 - (d) far from reactivity
6. _____ Likewise, the first column in Group 1, has characteristic number of 1. The second column in the group has characteristic number of
 - (a) 2
 - (b) 10
 - (c) 18
 - (d) 32
7. The groups in the periodic table are numbered 1 to 18 and are divided into (a) 10
 - (a) 10 s-block elements (including noble)
 - (b) 10 p-block elements (including noble)
 - (c) 10 d-block elements (including noble)
 - (d) 10 f-block elements (including noble)

Periodic Law Chapter 5 Test Answers

Robert M. Horowitz



Periodic Law Chapter 5 Test Answers:

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Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988

O'Sullivan & Hilliard's The Law of Contract Janet O'Sullivan, 2024 Shattered Symmetry Pieter Thyssen, Arnout Ceulemans, 2017-01-02 The standard model of subatomic particles and the periodic table of the atoms have the common goal to bring order in the bewildering chaos of the constituents of matter Their success relies on the presence of fundamental symmetries in their core The purpose of the book is to share the admiration for the power and the beauty of these symmetries The reader is taken on a journey from the basic geometric symmetry group of a circle to the sublime dynamic symmetries that govern the motions of the particles The trail follows the lines of parentage linking groups upstream to the unitary symmetry of the eightfold way of quarks and to the four dimensional symmetry of the hydrogen atom Along the way the theory of symmetry groups is gradually introduced with special emphasis on graphical representations The final challenge is to open up the structure of Mendeleev's table which goes beyond the symmetry of the hydrogen atom Breaking this symmetry to accommodate the multi electron atoms requires to leave the common ground of linear algebras and explore the potential of non linearity Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple

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