



Percent Yield

45 grams of nitrogen monoxide (NO) react with an excess of sodium hydroxide (NaOH) following the balanced equation above. 6 grams of water were obtained from the reaction

What is the % yield?

Nitrogen monoxide (NO) gmm:

$$(1)(14.0) + (1)(16.0) = \textcircled{30.0 \text{ g/mol}}$$

$$\text{Actual Yield} = 6 \text{ g H}_2\text{O}$$

Water (H₂O) gmm:

$$(2)(1.0) + (1)(16.0) = \textcircled{18.0 \text{ g/mol}}$$

$$\text{Theoretical Yield} = 6.75 \text{ g H}_2\text{O}$$

$$\frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100 = \textcircled{88.9\%}$$

$$\frac{45 \text{ g NO} \quad | \quad \cancel{1 \text{ mol NO}} \quad | \quad \cancel{1 \text{ mol H}_2\text{O}} \quad | \quad 18 \text{ g H}_2\text{O}}{30 \text{ g NO} \quad | \quad \cancel{4 \text{ mol NO}} \quad | \quad \cancel{1 \text{ mol H}_2\text{O}}} = 6.75 \text{ g H}_2\text{O}$$

Percent Yield Problems And Answers

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



Percent Yield Problems And Answers:

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