

A1 States of matter

1. a) i) Ammonium chloride (1)
 ii) Diffusion (1)
 Diffusion is the movement of particles from an area of higher concentration to an area of lower concentration until the particles are evenly distributed. (2)
 iii) The ammonia solution gave off ammonia gas and the hydrochloric acid gave off hydrogen chloride gas. The particles of the two gases diffused along the tube, collided and reacted to form ammonium chloride. (3)
- b) i)  (1)
 ii) The distilled water had a higher water content than the cytoplasm inside the paw paw cells and the cell membranes were differentially permeable. The water molecules were able to move through the membranes into the cells causing the cells to swell and the strips to increase in volume. (3)
 iii) Osmosis (1)
- c) Sodium chloride draws water out of the cells of the fish and any micro-organisms by osmosis. Water is, therefore, not available for chemical reactions in the cells, some of which cause decay and micro-organisms can't grow. (2)

2. a)

Property	Solid	Liquid	Gas
Volume		definite	variable; the volume is the same as the entire container
Arrangement of particles	packed closely together, usually in a regular way		randomly arranged with large spaces between
Energy of particles	have very small amounts of kinetic energy	have medium amounts of kinetic energy	

(6)

- b) i) The particles have large spaces between them, so they can be easily pushed closer together. (1)
 ii) The particles are packed tightly together with very little empty space between. (1)
 iii) The particles move around rapidly and have weak forces of attraction between them, so they spread out to fill any available space in the container. (2)

- c) i) A: Melting B: Boiling or evaporation (4)
 C: Freezing D: Condensation
 ii) Iodine or carbon dioxide or ammonium chloride (1)
 or naphthalene (1)
 d) i) Liquid (1)
 ii) 54 °C (1)
 iii)  (1)

A2 Mixtures and separations

1. a) i)

	Pure substance	Mixture
Composition		variable
Properties	fixed and constant	variable; the components retain their individual properties

 (3)
- ii) An element is a pure substance that cannot be broken down into any simpler substances by any ordinary chemical or physical means. A compound is a pure substance that contains two or more different types of element that are bonded together chemically in fixed proportions and in such a way that their properties have changed. (2)
- b) i) The particles in a suspension are larger than those in a colloid. (1)
 ii) The particles in a suspension settle if left undisturbed, whereas the particles in a colloid never settle. (1)
 iii) Light does not pass through a suspension, whereas most colloids scatter light. (1)
- c) i) A solution is a homogeneous mixture consisting of two or more components, one of which is usually a liquid. (1)
 ii) Solubility is the mass of solute that will saturate 100 g of solvent at a given temperature. (1)
- d) i) 17 g per 100 g water (1)
 ii) 45 °C (1)
 iii) At 76 °C, 54 g of Z saturate 100 g water
 At 36 °C, 12.5 g of Z saturate 100 g water
 $\therefore$ mass of Z crystallising out of a saturated solution containing 100 g water = $54 - 12.5$ g = 41.5 g (3)
- iv) At 62 °C, 39 g of Z saturate 100 g water
 $\therefore$ at 62 °C, 45 g of Z saturate $\frac{100}{39} \times 45$ g water = 115.4 g of water (2)
- v) At 55 °C, 33 g of Z saturate 100 g water
 $\therefore$ at 55 °C, $\frac{33}{100} \times 350$ g of Z saturate 350 g water = 115.5 g of Z (2)

Pearson Chemistry Workbook Answers States Of Matter

JR Anderson



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Table of Contents Pearson Chemistry Workbook Answers States Of Matter

1. Understanding the eBook Pearson Chemistry Workbook Answers States Of Matter
 - The Rise of Digital Reading Pearson Chemistry Workbook Answers States Of Matter
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Workbook Answers States Of Matter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Workbook Answers States Of Matter
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Workbook Answers States Of Matter
 - Personalized Recommendations

- Pearson Chemistry Workbook Answers States Of Matter User Reviews and Ratings
- Pearson Chemistry Workbook Answers States Of Matter and Bestseller Lists
- 5. Accessing Pearson Chemistry Workbook Answers States Of Matter Free and Paid eBooks
 - Pearson Chemistry Workbook Answers States Of Matter Public Domain eBooks
 - Pearson Chemistry Workbook Answers States Of Matter eBook Subscription Services
 - Pearson Chemistry Workbook Answers States Of Matter Budget-Friendly Options
- 6. Navigating Pearson Chemistry Workbook Answers States Of Matter eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry Workbook Answers States Of Matter Compatibility with Devices
 - Pearson Chemistry Workbook Answers States Of Matter Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Workbook Answers States Of Matter
 - Highlighting and Note-Taking Pearson Chemistry Workbook Answers States Of Matter
 - Interactive Elements Pearson Chemistry Workbook Answers States Of Matter
- 8. Staying Engaged with Pearson Chemistry Workbook Answers States Of Matter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Workbook Answers States Of Matter
- 9. Balancing eBooks and Physical Books Pearson Chemistry Workbook Answers States Of Matter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Workbook Answers States Of Matter
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Workbook Answers States Of Matter
 - Setting Reading Goals Pearson Chemistry Workbook Answers States Of Matter
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Workbook Answers States Of Matter
 - Fact-Checking eBook Content of Pearson Chemistry Workbook Answers States Of Matter

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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