- Pearsons Learning: Chapter 14: Solutions
- What is the volume/volume percent of a solution prepared from 12.3 mL ethanol in 150.0 mL water? Assume volumes are additive.
- Answer: percent concentration = amount of solute x100 amount of solute + omount of solvent percent concentration (ethanol) = 150.ml ethanol x 100 = 92.9% 100% - 92.4% = 7.58% (v/v) ethanol
- Which expression is equivalent to 5.45% (v/v) ethanol?
- Answer: percent concentration = amount of solute x100
- 5.45% (v/v) ethanul = 100 = 0.0545 ml 0.0545 ml = 5.45 ml ethanol
- What mass of NaCl is present in 50.0 g of 3.75% (m/m) NaCl solution?
- Answer: 3.75% (m/m) Nac/ + 100 = 0.03750 mass percent concentration = mass of solute x 100

  96. 376% = 50.09 mass of solute + mass of solvent 50.0g + 1.88g 100% - 96.376% = 3.624 % (m/m) Nacl
- Mass of Nacl = 1.889
- What is the osmotic pressure of a 0.850 M solution of glucose in water at 35 °C?

# **Pearson Chapter 14 Mixture And Solutions Worksheet**

**David B. Troy, Paul Beringer** 

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