

$$2. [\text{OH}^-] = 1.0 \times 10^{-14} \text{ M}$$

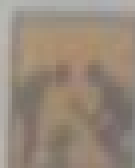
$$[\text{H}^+] = ?$$

$$K_w = [\text{H}^+][\text{OH}^-]$$

$$1.0 \times 10^{-14} = [\text{H}^+](1.0 \times 10^{-14})$$

$$[\text{H}^+] = 1.0 \times 10^{-14} \text{ M}$$

Because $1.0 \times 10^{-14} \text{ M} < 1.0 \times 10^{-7} \text{ M}$, the solution is acidic.



Self Assessment Exercise 10.3

1. A soft drink has $[\text{H}^+] = 3 \times 10^{-4} \text{ M}$. Is the drink acidic, basic, or neutral?
2. Ordinary vinegar is approximately 1M CH_3COOH . Concentration of H^+ is $4.2 \times 10^{-3} \text{ M}$. Is vinegar acidic, basic, or neutral?
3. A student measures the $[\text{OH}^-]$ of a solution and obtains a value of $1.0 \times 10^{-10} \text{ M}$. Is the solution acidic, basic, or neutral?

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Karin Nielsen-Saines



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