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 $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$ total ionic $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$ net ionic $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water
 $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ the spectator ions in $\text{KOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$ are $\text{K}^+(\text{aq})$ and $\text{NO}_3^-(\text{aq})$ how to identify and list the

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 $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{KCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{KNO}_3(\text{aq})$ complete ionic equation an equation written with all soluble

strong electrolytes shown as ions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{Cl}^- + \text{Cu}^{2+} \rightarrow \text{CuCl}_2 \text{ s}$ $2\text{Cl}^- + \text{Cu}^{2+} \rightarrow \text{CuCl}_2 \text{ s}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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