

Calculus-Based Properties (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 5.56 \text{ mol/kg} = 10.3^\circ C \quad \Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 5.56 \text{ mol/kg}$$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of NaCl and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$, $K_f(\text{NaCl}) = 3.72^\circ C/m$, $K_b(\text{NaCl}) = 1.02^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{0.100 \text{ kg}} = 1.71 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 3.72^\circ C/m \times 1.71 \text{ mol/kg} = 6.36^\circ C \quad \Delta T_b = K_b \cdot m = 1.02^\circ C/m \times 1.71 \text{ mol/kg}$$

3. How many grams of $C_6H_{12}O_6$ must be added to 100 g of H_2O to raise the boiling-point of the solution to $100.0^\circ C$? Assume that the solute does not dissociate. Equations: $K_b = 0.512^\circ C/m$.

$$\Delta T_b = K_b \cdot m \quad 0.512^\circ C/m \times m = 100.0^\circ C - 100.0^\circ C = 0.0^\circ C$$

$$m = \frac{\Delta T_b}{K_b} = \frac{0.0^\circ C}{0.512^\circ C/m} = 0 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$0 \text{ mol} = 0.100 \text{ kg} \times m$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$0 \text{ mol} = 0.100 \text{ kg} \times m \quad 0 \text{ mol} = 0.100 \text{ kg} \times m$$

$$0 \text{ mol} = 0.100 \text{ kg} \times m \quad 0 \text{ mol} = 0.100 \text{ kg} \times m$$

$$\text{moles solute} = 0 \text{ mol} \quad \text{grams solute} = 0 \text{ g}$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 100 g of water is prepared. The freezing-point of this solution is $-0.5^\circ C$. The normal freezing-point of pure water is $0.0^\circ C$. What is the boiling-point of this solution? Assume that the solute does not dissociate. Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\Delta T_f = K_f \cdot m \quad 0.5^\circ C = 1.86^\circ C/m \times m \quad m = \frac{0.5^\circ C}{1.86^\circ C/m} = 0.27 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.27 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$0.27 \text{ mol} = 0.100 \text{ kg} \times m \quad 0.27 \text{ mol} = 0.100 \text{ kg} \times m$$

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$$\Delta T_b = K_b \cdot m \quad 0.512^\circ C/m \times 0.27 \text{ mol/kg} = 0.14^\circ C$$

$$\text{Boiling-point} = 100.0^\circ C + 0.14^\circ C = 100.14^\circ C$$

Rosengarten Colligative Properties Homework Answers

Kathleen Armour



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