

Naming Molecular Compounds

How are the chemical formula and name of a molecular compound related?

Why?

When you began chemistry class this year, you probably already knew that the chemical formula for carbon dioxide was CO_2 . Today you will find out why CO_2 is named that way. Naming chemical compounds correctly is of paramount importance. The slight difference between the names carbon monoxide (CO , a poisonous, deadly gas) and carbon dioxide (CO_2 , a greenhouse gas that we exhale when we breathe out) can be the difference between life and death! In this activity you will learn the naming system for molecular compounds.

Model 1 – Molecular Compounds

Molecular Formula	Number of Atoms of First	Number of Atoms of Second	Name of Compound
ClF	1	1	Chlorine monofluoride
ClF_5	1	5	Chlorine pentafluoride
CO	1	1	Carbon monoxide
CO_2	1	5	Carbon dioxide
Cl_2O	2	1	Dichlorine monoxide
PCl_5	1	5	Phosphorus pentachloride
N_2O_5	2	5	Dinitrogen pentoxide

- Fill in the table to indicate the number of atoms of each type in the molecular formula.
- Examine the molecular formulas given in Model 1 for various molecular compounds.
 - How many different *elements* are present in each compound shown?
Two
 - Do the compounds combine metals with metals, metals with nonmetals, or nonmetals with nonmetals?
Non-metals with non-metals
 - Based on your answer to b, what type of bonding must be involved in molecular compounds?
Covalent bonds
- Find all of the compounds in Model 1 that have chlorine and fluorine in them. Explain why the name "chlorine fluoride" is not sufficient to identify a specific compound.
The number of fluorine atoms is different in the two compounds, but name does not reflect the difference
- Assuming that the name of the compound gives a clue to its molecular formula, predict how many atoms each of these prefixes indicates, and provide two examples.

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L Towner

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Table of Contents Pogil Chemistry Answer Key Naming Molecular Compounds

1. Understanding the eBook Pogil Chemistry Answer Key Naming Molecular Compounds
 - The Rise of Digital Reading Pogil Chemistry Answer Key Naming Molecular Compounds
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Chemistry Answer Key Naming Molecular Compounds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Chemistry Answer Key Naming Molecular Compounds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Chemistry Answer Key Naming Molecular Compounds

- Personalized Recommendations
- Pogil Chemistry Answer Key Naming Molecular Compounds User Reviews and Ratings
- Pogil Chemistry Answer Key Naming Molecular Compounds and Bestseller Lists
- 5. Accessing Pogil Chemistry Answer Key Naming Molecular Compounds Free and Paid eBooks
 - Pogil Chemistry Answer Key Naming Molecular Compounds Public Domain eBooks
 - Pogil Chemistry Answer Key Naming Molecular Compounds eBook Subscription Services
 - Pogil Chemistry Answer Key Naming Molecular Compounds Budget-Friendly Options
- 6. Navigating Pogil Chemistry Answer Key Naming Molecular Compounds eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Chemistry Answer Key Naming Molecular Compounds Compatibility with Devices
 - Pogil Chemistry Answer Key Naming Molecular Compounds Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Chemistry Answer Key Naming Molecular Compounds
 - Highlighting and Note-Taking Pogil Chemistry Answer Key Naming Molecular Compounds
 - Interactive Elements Pogil Chemistry Answer Key Naming Molecular Compounds
- 8. Staying Engaged with Pogil Chemistry Answer Key Naming Molecular Compounds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Chemistry Answer Key Naming Molecular Compounds
- 9. Balancing eBooks and Physical Books Pogil Chemistry Answer Key Naming Molecular Compounds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Chemistry Answer Key Naming Molecular Compounds
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Chemistry Answer Key Naming Molecular Compounds
 - Setting Reading Goals Pogil Chemistry Answer Key Naming Molecular Compounds
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Chemistry Answer Key Naming Molecular Compounds

- Fact-Checking eBook Content of Pogil Chemistry Answer Key Naming Molecular Compounds
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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