

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in organisms, it is important to understand the ability of buffers to resist change. This property is called **buffering capacity** and is usually expressed as millimoles per liter of buffering agent or as millimoles per liter of acid or base that can be added to 1 L of solution before the pH of the solution changes by one unit. Buffering capacity is the ability of a buffer to resist change in pH.

Table 1 – Ion Concentrations for Acids and Bases

Number	Substance	Acidity, Alkalinity or Neutrality	$[H^+]$ (M)	$[OH^-]$ (M)	pH or pOH
1	0.1 M HCl (strong acid)	Acidic	1.0×10^{-1} M	1.0×10^{-13} M	1.0
2	0.001 M HCl (strong acid)	Acidic	1.0×10^{-3} M	1.0×10^{-11} M	3.0
3	0.00001 M HCl (strong acid)	Acidic	1.0×10^{-5} M	1.0×10^{-9} M	5.0
4	0.0000001 M HCl (strong acid)	Acidic	1.0×10^{-7} M	1.0×10^{-7} M	7.0
5	0.00000001 M HCl (strong acid)	Acidic	1.0×10^{-8} M	1.0×10^{-6} M	8.0
6	0.000000001 M HCl (strong acid)	Acidic	1.0×10^{-9} M	1.0×10^{-5} M	9.0
7	0.0000000001 M HCl (strong acid)	Acidic	1.0×10^{-10} M	1.0×10^{-4} M	10.0
8	0.00000000001 M HCl (strong acid)	Acidic	1.0×10^{-11} M	1.0×10^{-3} M	11.0
9	0.000000000001 M HCl (strong acid)	Acidic	1.0×10^{-12} M	1.0×10^{-2} M	12.0
10	0.0000000000001 M HCl (strong acid)	Acidic	1.0×10^{-13} M	1.0×10^{-1} M	13.0

1. What does the number 13.0 pH indicate? (13.0 is the pH)

It indicates the concentration of H^+ ions.

2. In Solution 1, which one has a higher concentration, hydrochloric acid or hydroxide ions?

Hydrochloric acid.

3. How did you determine the concentration of hydrochloric acid ions and hydroxide ions? (Use Solution 1 as Model 1.)

From the pH value, I determined the concentration of H^+ ions. I then used the K_w value to determine the concentration of OH^- ions.

4. How did you determine the concentration of hydroxide ions? (Use Solution 1 as Model 1.)

From the pH value, I determined the concentration of H^+ ions. I then used the K_w value to determine the concentration of OH^- ions.

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