

The image displays a standard periodic table of elements, color-coded by groups. The title "PERIODIC TABLE OF ELEMENTS" is centered at the top. The table is organized into rows and columns, with element symbols and names provided for each. The groups are labeled as follows:

- 1A: Hydrogen (H)
- 2A: Helium (He)
- 3A: Boron (B)
- 4A: Carbon (C)
- 5A: Nitrogen (N)
- 6A: Oxygen (O)
- 7A: Fluorine (F)
- 8A: Neon (Ne)
- 9A: Sodium (Na)
- 10A: Magnesium (Mg)
- 11A: Aluminum (Al)
- 12A: Silicon (Si)
- 13A: Phosphorus (P)
- 14A: Sulfur (S)
- 15A: Chlorine (Cl)
- 16A: Argon (Ar)
- 17A: Potassium (K)
- 18A: Calcium (Ca)
- 19A: Scandium (Sc)
- 20A: Titanium (Ti)
- 21A: Vanadium (V)
- 22A: Chromium (Cr)
- 23A: Manganese (Mn)
- 24A: Iron (Fe)
- 25A: Cobalt (Co)
- 26A: Nickel (Ni)
- 27A: Copper (Cu)
- 28A: Zinc (Zn)
- 29A: Gallium (Ga)
- 30A: Germanium (Ge)
- 31A: Arsenic (As)
- 32A: Selenium (Se)
- 33A: Bromine (Br)
- 34A: Krypton (Kr)
- 35A: Rubidium (Rb)
- 36A: Strontium (Sr)
- 37A: Yttrium (Y)
- 38A: Zirconium (Zr)
- 39A: Niobium (Nb)
- 40A: Molybdenum (Mo)
- 41A: Technetium (Tc)
- 42A: Ruthenium (Ru)
- 43A: Rhodium (Rh)
- 44A: Palladium (Pd)
- 45A: Silver (Ag)
- 46A: Cadmium (Cd)
- 47A: Indium (In)
- 48A: Tin (Sn)
- 49A: Antimony (Sb)
- 50A: Tellurium (Te)
- 51A: Iodine (I)
- 52A: Xenon (Xe)
- 53A: Cesium (Cs)
- 54A: Barium (Ba)
- 55A: Lanthanum (La)
- 56A: Hafnium (Hf)
- 57A: Tantalum (Ta)
- 58A: Tungsten (W)
- 59A: Rhenium (Re)
- 60A: Osmium (Os)
- 61A: Iridium (Ir)
- 62A: Platinum (Pt)
- 63A: Gold (Au)
- 64A: Mercury (Hg)
- 65A: Thallium (Tl)
- 66A: Lead (Pb)
- 67A: Bismuth (Bi)
- 68A: Polonium (Po)
- 69A: Astatine (At)
- 70A: Radon (Rn)
- 71A: Francium (Fr)
- 72A: Radium (Ra)
- 73A: Actinium (Ac)
- 74A: Rutherfordium (Rf)
- 75A: Dubnium (Db)
- 76A: Seaborgium (Sg)
- 77A: Bohrium (Bh)
- 78A: Hassium (Hs)
- 79A: Meitnerium (Mt)
- 80A: Ds
- 81A: Rg
- 82A: Cn
- 83A: Nh
- 84A: Fl
- 85A: Mc
- 86A: Lv
- 87A: Ts
- 88A: Og

The Lanthanides and Actinides are shown separately at the bottom of the table, color-coded by group.

LANTHANIDES

- La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu

ACTINIDES

- Ac, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr



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Mendeleev's most important writings on the periodic law. Thirteen papers and essays divided into three groups reflect the period corresponding to the initial establishment of the periodic law (three papers, 1869–71), a period of priority disputes and experimental confirmations (five papers, 1871–86), and a final period of general acceptance for the law and increasing international recognition for Mendeleev (five papers, 1887–1905). A single, easily accessible source for Mendeleev's principle papers, this volume offers a history of the development of the periodic law written by the law's own founder.

Periodic Table Advanced BarCharts, Inc., Mark Jackson, 2014-12-31. The ultimate reference tool and lab partner for any student of science. Durably laminated and designed to fit as much info as possible in this handy 6-page format. Separate property tables are broken out for the ease of locating trends while studying and working. Other pages offer essential notes about the table's organization and history. Consistently a best seller since its first creation.

Periodic Table Advanced BarCharts, Inc., 2014.

150 Years of the Periodic Table Carmen J. Giunta, Vera V. Mainz, Gregory S. Girolami, 2021-07-04. This book provides an overview of the origins and evolution of the periodic system from its prehistory to the latest synthetic elements and possible future additions. The periodic system of the elements first emerged as a comprehensive classificatory and predictive tool for chemistry during the 1860s. Its subsequent embodiment in various versions has made it one of the most recognizable icons of science. Based primarily on a symposium titled 150 Years of the Periodic Table and held at the August 2019 national meeting of the American Chemical Society, this book describes the origins of the periodic law, developments that led to its acceptance, chemical families that the system struggled to accommodate, extension of the periodic system to include synthetic elements, and various cultural aspects of the system that were celebrated during the International Year of the Periodic Table.

How the World Works: The Periodic Table Anne Rooney, 2019-02. Everything in the universe is made of chemical elements, including you. In 1869, Russian chemist Dmitri Mendeleev produced a periodic table designed to illustrate the properties of the known elements. This arrangement of the elements in order of increasing atomic number was an important milestone in the development of chemistry and led to the establishment of periodic law. Written in a straightforward, easily comprehensible way, *The Periodic Table* explores the story of each element, describing the people who discovered them and taking us on a journey of discovery into what the whole world is made of.

The Periodic Table D. H. Rouvray, R. Bruce King, 2004. This book starts with chapters that trace the early history and development of the Periodic Table. The subsequent development of the Table is then presented in chapters that discuss the structure and characteristics of the Table, probe its group-theoretical and quantum-theoretical basis, examine its foundations, and explore its many uses and applications.

The Periodic Table II D. Michael P. Mingos, 2020-02-05. As 2019 has been declared the International Year of the Periodic Table, it is appropriate that Structure and Bonding marks this anniversary with two special volumes. In 1869, Dmitri Ivanovitch Mendeleev first proposed his periodic table of the elements. He is given the major credit for proposing the conceptual framework used by chemists to systematically interrelate the chemical properties of the elements. However,

the concept of periodicity evolved in distinct stages and was the culmination of work by other chemists over several decades. For example, Newland's Law of Octaves marked an important step in the evolution of the periodic system since it represented the first clear statement that the properties of the elements repeated after intervals of 8. Mendeleev's predictions demonstrated in an impressive manner how the periodic table could be used to predict the occurrence and properties of new elements. Not all of his many predictions proved to be valid, but the discovery of scandium, gallium, and germanium represented sufficient vindication of its utility and they cemented its enduring influence. Mendeleev's periodic table was based on the atomic weights of the elements and it was another 50 years before Moseley established that it was the atomic number of the elements that was the fundamental parameter and this led to the prediction of further elements. Some have suggested that the periodic table is one of the most fruitful ideas in modern science and that it is comparable to Darwin's theory of evolution by natural selection proposed at approximately the same time. There is no doubt that the periodic table occupies a central position in chemistry. In its modern form, it is reproduced in most undergraduate inorganic textbooks and is present in almost every chemistry lecture room and classroom. This second volume provides chemists with an overview of the important role played by the Periodic Table in advancing our knowledge of solid state and bioinorganic chemistry. It also illustrates how it has been used to fine-tune the properties of compounds which have found commercial applications in catalysis, electronics, ceramics, and in medicinal chemistry.

The Periodic Table Eric R. Scerri, 2019-07-25

The periodic table of elements first encountered by many of us at school provides an arrangement of the chemical elements ordered by their atomic number, electron configuration, and recurring chemical properties, and divided into periodic trends. In this Very Short Introduction, Eric R. Scerri looks at the trends in properties of elements that led to the construction of the table and shows how the deeper meaning of the table's structure gradually became apparent with the development of atomic theory and, in particular, quantum mechanics, which underlies the behaviour of all of the elements and their compounds. This new edition, publishing in the International Year of the Periodic Table, celebrates the completion of the seventh period of the table with the ratification and naming of elements 113, 115, 117, and 118 as nihonium, moscovium, tennessine, and oganesson. Eric R. Scerri also incorporates new material on recent advances in our understanding of the origin of the elements, as well as developments concerning group three of the periodic table.

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