

Percent Yield

1. In a certain chemical reaction, 1.00 g of reactant A (molar mass 100.0 g/mol) reacts with excess of reactant B (molar mass 50.0 g/mol) to produce 0.50 g of product C (molar mass 150.0 g/mol). What is the percent yield of product C?
2. A certain chemical reaction produces 10.0 g of product C (molar mass 100.0 g/mol) from 5.0 g of reactant A (molar mass 50.0 g/mol) and 10.0 g of reactant B (molar mass 100.0 g/mol). What is the percent yield of product C?
3. A certain chemical reaction produces 10.0 g of product C (molar mass 100.0 g/mol) from 5.0 g of reactant A (molar mass 50.0 g/mol) and 10.0 g of reactant B (molar mass 100.0 g/mol). What is the percent yield of product C?
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Percent Yield Skill Practice 44

LL Leslie



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