

CHAPTER 15 STUDY GUIDE

Energy and Chemical Change

Section 15.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- | | |
|-------|---|
| _____ | 1. <i>Energy</i> is the ability to do work or produce heat. |
| _____ | 2. The law of conservation of energy states that energy <i>can be</i> created and destroyed. |
| _____ | 3. Chemical potential energy is energy stored in a substance because of its <i>composition</i> . |
| _____ | 4. <i>Heat</i> is a form of energy that flows from a warmer object to a cooler object. |
| _____ | 5. A calorie is the amount of energy required to raise the temperature of <i>one gram</i> of pure water by one degree Celsius. |
| _____ | 6. A <i>calorie</i> is the SI unit of heat and energy. |
| _____ | 7. The <i>specific heat</i> of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius. |
| _____ | 8. <i>Kinetic energy</i> is energy of motion. |
| _____ | 9. Chemicals participating in a chemical reaction contain only <i>potential energy</i> . |
| _____ | 10. One nutritional Calorie is equal to <i>100 calories</i> . |
| _____ | 11. One calorie equals <i>4.184 joules</i> . |
| _____ | 12. When a fuel is burned, some of its <i>chemical potential energy</i> is lost as heat. |
| _____ | 13. To convert kilojoules to joules, <i>divide</i> the number of kilojoules by 1000 joules/1 kilojoule. |

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

Pearson Chemistry Ch 15 Answer Key

Thomas Edward Thorpe



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Pearson Edexcel A Level Chemistry (Year 1 and Year 2) Andrew Hunt, Graham Curtis, Graham Hill, 2019-07-15 Develop and assess your students knowledge and skills throughout A level with worked examples practical assessment guidance and differentiated end of topic questions in this updated all in one textbook for Years 1 and 2 Combining everything your students need to know for the Pearson Edexcel A level Chemistry specification this revised textbook will Identify the level of your students understanding with diagnostic questions and a summary of prior knowledge at the start of the Student Book Provide support for all 16 required practicals with various activities and questions along with a Practical chapter covering procedural understanding and key ideas related to measurement Improve mathematical skills with plenty of worked examples including notes on methods to help explain the strategies for solving each type of problem Offer plenty of practice with Test yourself questions to help students assess their understanding and measure progress Encourage further reading and study with short passages of extension material Develop understanding with free online access to Test yourself answers and an extended glossary Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation Chemistry 'O' Level Rex M. Heyworth, 2007 Characterization of a Biodegradable, PH-sensitive Chitosan-PVA Hydrogel Tao Wang, 2002 *Chemistry* McMurry, Pearson, 2003-07 *A Dictionary of Applied Chemistry* Thomas Edward Thorpe, 1913 **Remington** David B. Troy, Paul Beringer, 2006 For over 100 years Remington has been the definitive textbook and reference on the science and practice of pharmacy This Twenty First Edition keeps pace with recent changes in the pharmacy curriculum and professional pharmacy practice More than 95 new contributors and 5 new section editors provide fresh perspectives on the field New chapters include pharmacogenomics application of ethical principles to practice dilemmas technology and automation professional communication medication errors re engineering pharmacy practice management of special risk medicines specialization in pharmacy practice disease state management emergency patient care and wound care Purchasers of this textbook are entitled to a new fully indexed Bonus CD ROM affording instant access to the full content of Remington in a convenient and portable format **Chemistry O Level Mauritius Edition** Pearson Education, Limited, 2007-06-26

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Introductory Nanoscience Masaru Kuno, 2011-08-19 Designed for students at the senior undergraduate and first year graduate level Introductory Nanoscience takes a quantitative approach to describing the physical and chemical principles behind what makes nanostructures so fascinating This textbook provides a foundation for understanding how properties of materials change when scaled to nano size explaining how we may predict behavior and functionality Stoichiometry and Research Alessio Innocenti, 2012-03-07 The aim of this book is to provide an overview of the importance of stoichiometry in the biomedical field It proposes a collection of selected research articles and reviews which provide up to date information related to stoichiometry at various levels The first section deals with host guest chemistry focusing on selected calixarenes cyclodextrins and crown ethers derivatives In the second and third sections the book presents some issues concerning stoichiometry of metal complexes and lipids and polymers architecture The fourth section aims to clarify the role of stoichiometry in the determination of protein interactions while in the fifth section some selected experimental techniques applied to specific systems are introduced The last section of the book is an attempt at showing some interesting connections between biomedicine and the environment introducing the concept of biological stoichiometry On this basis the present volume would definitely be an ideal source of scientific information to researchers and scientists involved in biomedicine biochemistry and other areas involving stoichiometry evaluation Lectures on Physics Richard P. Feynman, 2006 Study Guide Essential Biology with Physiology Edward J. Zalisko, 2003-07 Students can master key concepts and earn a better grade with the thought provoking exercises found in this study guide Study advice tables quizzes and crossword puzzles help students test their understanding of biology The Study Guide also includes references to student media activities on the

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Physical Chemistry Robert J. Silbey, Robert A. Alberty, George A. Papadantonakis, Mouni G. Bawendi, 2022-06-15 Ever since Physical Chemistry was first published in 1913 it has remained a highly effective and relevant learning tool thanks to the efforts of physical chemists from all over the world Each new edition has benefited from their suggestions and expert advice The result of this remarkable tradition is now in your hands *A Dictionary of applied chemistry v. 5, 1913* Sir Thomas Edward Thorpe, 1913

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