# Chemical Reactions

the have discussed a large number of chemical processes. Some are accompanied by a change in colour, some by the coolution of a gas, some by the coolution of a gas, some by the coolution of a perceptiate, and so on. At the incidental level, different types of changes over. In some reactions, smaller parts add over. In some reactions, whereas in some others, bigger entities break into smaller others, one element displaces another from a compound whereas in some others, compounds exchange radicals among themselves. For a systematic study, therefore, it is essential that we classify the reactions into different types.

However, we will discuss first the change in energy in chemical reactions, which is common to all types.

# Change in Energy in a Reaction

All reactions are accompanied by a change in energy. Energy (in the form of heat or light) is either given out or taken in during a chemical reaction.

A chemical reaction in which heat or light is evolved is called an exciberanic reaction.

A chemical reaction in which heat or light-in absorbed is called an endothermic reaction.

In an exothermic process, heat or light is given out to the surroundings. That is why you feel hot when you accidentally touch anything burning. On the other hand, in an endothermic process, heat or light is taken

in it absorbed from the surrounding. It is why a cuite of ice is cold to touch. The is why a cuite of ice is cold to touch. The itends to melt when you touch it. The itends to melt when you touch it. The itends to melt when you reaction and itended in the interest of ice is a physical change, not a charge of ice is a physical change, not a charge reaction.

# Why a Change in Energy in a Reaction

As we know, any two atoms in a molecule, held together by a force of attraction could be chemical bond. And also that a chemical he is much stronger than an intermolecular is We also know that it is the atoms that to part in a chemical reaction.

Thus, we can consider a chemical mento be a result of the following phenomen

- molecules so as to set the atoms in a new combination (Energy is about in the process.)
- (ii) The forming of fresh bonds benshing the new partners so that the processing out in the processing out in the processing the processing out in the processing the proce

The energy required for the breaking of the bonds of the reactant molecules is not same as that given out in the formation of bonds between the new partners. When energy given out in the formation of the breaking of the bonds, the surplus energy is evolved and process is exothermic.

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**Stephen J. Bonsor, Gavin Pearson** 

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