

## Topic 20 Organic synthesis

#### 20A Organic structures

#### 20A.1 Organic analysis

- T does not contain a C=C (no immediate decolourisation of bromine water).

  T does not contain a -COOH group (does not react with aqueous sodium carbonate). T contains one -OH group (produces 1 mol of HCl per mole of compound). T contains a carbonyl group (orange precipitate with 2,4-DNPH). The possible structural formulae are:

  CH<sub>3</sub>COCH<sub>2</sub>OH CH<sub>3</sub>CHOHCHO CH<sub>2</sub>OHCH<sub>2</sub>CHO

  Isomer A Isomer B Isomer C
  - (b) Relative molecular mass of T is 74.
    A m/z value of 29 suggests a CH<sub>3</sub>CH<sub>2</sub>— group or a —CHO group, so T cannot be isomer A. A m/z value of 59 (74 15) suggests a CH<sub>3</sub>— group, so T cannot be isomer C.
    T is isomer B = CH<sub>3</sub>CHOHCHO
- C (\$2.17 + 12.0) = 4.35

  H (13.04 + 1.0) = 13.04

  O (\$4.78 + 16.0) = 2.17

  Dividing by 2.17 gives the whole number ratio of C:H:O of 2:6:1

  The empirical formula is C<sub>2</sub>H<sub>6</sub>O

  The empirical formula mass = (12.0 × 2) + (1.0 × 6) + (16.0 × 1) = 46

  The empirical formula mass = the relative molecular mass, so the molecular formula is C<sub>2</sub>H<sub>6</sub>O

  The peak at chemical shift 1.2 ppm is caused by the protons in a CH<sub>3</sub> group in an alkane chain. It is a triplet, so it is attached to a CH<sub>2</sub> group.

  The peak at 3.7 is caused by the protons in a CH<sub>2</sub> group attached to an oxygen atom. It is a quartet, so it is also attached to a CH<sub>3</sub> group.

  The peak at 2.6 is caused by the protons in an OH group.

  The structural formula of the compound is CH<sub>3</sub>CH<sub>2</sub>OH.

#### 20A.2 Organic synthesis

$$\begin{array}{c} CH_3 \\ -CH_3 \\ CH_3 \\ -CH_2 \\ -CH_3 \\ CH_3 \end{array} \begin{array}{c} CH -CH_2 -CH_3 \\ -CH_3 \\ -CH_3 \\ -CH_3 \end{array}$$

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**S Ben Porath** 

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#### **Table of Contents Pearson Education Topic 20 6 Answers**

- 1. Understanding the eBook Pearson Education Topic 20 6 Answers
  - The Rise of Digital Reading Pearson Education Topic 20 6 Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Pearson Education Topic 20 6 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Education Topic 20 6 Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pearson Education Topic 20 6 Answers
  - Personalized Recommendations
  - Pearson Education Topic 20 6 Answers User Reviews and Ratings
  - Pearson Education Topic 20 6 Answers and Bestseller Lists
- 5. Accessing Pearson Education Topic 20 6 Answers Free and Paid eBooks
  - Pearson Education Topic 20 6 Answers Public Domain eBooks
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  - Pearson Education Topic 20 6 Answers Budget-Friendly Options

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  - Pearson Education Topic 20 6 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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  - Highlighting and Note-Taking Pearson Education Topic 20 6 Answers
  - Interactive Elements Pearson Education Topic 20 6 Answers
- 8. Staying Engaged with Pearson Education Topic 20 6 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Education Topic 20 6 Answers
- 9. Balancing eBooks and Physical Books Pearson Education Topic 20 6 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Education Topic 20 6 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Education Topic 20 6 Answers
  - Setting Reading Goals Pearson Education Topic 20 6 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Education Topic 20 6 Answers
  - Fact-Checking eBook Content of Pearson Education Topic 20 6 Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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