



COVALENT BONDING

Practice Problems

Do your answers to these very carefully.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following as ionic or molecular.

- a. NaCl (ionic) b. H_2O (molecular) c. HCl (molecular)
 d. KNO_3 (ionic) e. H_2SO_4 (molecular)

2. Which of the following is a molecular compound?

- a. H_2O b. NaCl c. KNO_3
 d. H_2SO_4 e. H_2SO_3

3. Which of the following is a molecular compound? (Note: H_2O is a molecular compound.)

a. H_2O (molecular) b. H_2SO_4 (molecular) c. H_2SO_3 (molecular) d. H_2SO_2 (molecular) e. H_2SO (molecular)

4. Which of the following is a molecular compound? (Note: H_2O is a molecular compound.)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for H-F .



2. Draw the Lewis structure for CH_4 .



3. Draw the Lewis structure for CH_3Cl .



4. Draw the Lewis structure for H_2O .

5. Draw the Lewis structure for H_2O_2 .



6. Draw the Lewis structure for H_2O_2 .

SECTION 8.3 BONDING: THE ORBITAL

1. Which of the following is a bond formed by the overlap of two sp^3 hybrid orbitals?

a. H_2 (bond formed by the overlap of two 1s orbitals) b. H_2O (bond formed by the overlap of two sp^3 hybrid orbitals)

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Tomoaki Tanase, Yoichi Ishii



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The Proton in Chemistry R.P. Bell,2013-11-11 The first edition of this book was based on the lectures which I gave at Cornell University during 1958 as George Fisher Baker Lecturer and I would like to repeat my warmest thanks to Professor F A Long and the other members of the Department of Chemistry for their kindness and helpful advice The present edition was largely written during the tenure of a Visiting Professorship awarded by the Royal Society and the Israeli Academy of Sciences I am deeply indebted to both of these bodies and also to the hospitality of the Weizmann Institute of Science in particular to Professor David Samuel and Professor F S Klein of the Isotopes Research Department The subject as a whole has expanded greatly since 1959 especially in two fields namely the direct study of fast proton transfer reactions notably by the relaxation methods pioneered by Eigen and the experimental and theoretical study of hydrogen isotope effects In order to keep the size of the book within reasonable bounds it has been necessary to adopt a selective policy and this is particularly the case in Chapter 9 where I have chosen to treat a few types of reaction in some detail rather than to attempt a more complete coverage

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1. Understanding the eBook Pearson Chemistry 8 Covalent Bonding Answer Key
 - The Rise of Digital Reading Pearson Chemistry 8 Covalent Bonding Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry 8 Covalent Bonding Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Pearson Chemistry 8 Covalent Bonding Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry 8 Covalent Bonding Answer Key
 - Personalized Recommendations
 - Pearson Chemistry 8 Covalent Bonding Answer Key User Reviews and Ratings
 - Pearson Chemistry 8 Covalent Bonding Answer Key and Bestseller Lists

5. Accessing Pearson Chemistry 8 Covalent Bonding Answer Key Free and Paid eBooks
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 - Pearson Chemistry 8 Covalent Bonding Answer Key eBook Subscription Services
 - Pearson Chemistry 8 Covalent Bonding Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Pearson Chemistry 8 Covalent Bonding Answer Key Compatibility with Devices
 - Pearson Chemistry 8 Covalent Bonding Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry 8 Covalent Bonding Answer Key
 - Highlighting and Note-Taking Pearson Chemistry 8 Covalent Bonding Answer Key
 - Interactive Elements Pearson Chemistry 8 Covalent Bonding Answer Key
8. Staying Engaged with Pearson Chemistry 8 Covalent Bonding Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry 8 Covalent Bonding Answer Key
9. Balancing eBooks and Physical Books Pearson Chemistry 8 Covalent Bonding Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry 8 Covalent Bonding Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry 8 Covalent Bonding Answer Key
 - Setting Reading Goals Pearson Chemistry 8 Covalent Bonding Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry 8 Covalent Bonding Answer Key
 - Fact-Checking eBook Content of Pearson Chemistry 8 Covalent Bonding Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

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