

Chemistry I: Periodicity Worksheet

- 1) Arrange the following elements in order of increasing size: Al, B, C, K, and Na.

C	B	Al	Na	K
smallest				largest

- 2) Arrange the following elements in order of increasing size: Ca, Rb, P, Ge, and Sr.

P	Ge	Ca	Sr	Rb
smallest				largest

- 3) Circle the atom or ion in each pair that has the larger radius.

- a) Cl or Cl
b) Al or O
c) In or I

- 4) Select the atom or ion in each pair that has the larger radius.

- a) Cs or Rb
b) O^{2-} or O
c) Hg or As

- 5) Which of the following groups of elements is arranged correctly in order of increasing ionization energy?
(Answer by circling the correct answer.)

- a) $C < Si < Li < Ne$
 b) $Ne < Si < C < Li$
 c) $Li < Si < C < Ne$
 d) $Ne < C < Si < Li$

- 6) Arrange the following atoms in order of increasing ionization energy: Li, K, C, and N.

K Li C N
lowest highest

- 7) Arrange the following atoms in order of increasing ionization energy: Si, K, As, and Ca.

K	Ca	As	Si
lowest			highest

- 8) Compare the elements Na, Mg, O, and P.

- Which has the largest atomic radius? **Na**
- Which has the highest electron affinity? **O**
- Place the elements in order of increasing ionization energy.

lowest	Na	Mg	P	O	highest
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Periodicity Worksheet Answer Key

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Table of Contents Periodicity Worksheet Answer Key

1. Understanding the eBook Periodicity Worksheet Answer Key
 - The Rise of Digital Reading Periodicity Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodicity Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodicity Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodicity Worksheet Answer Key
 - Personalized Recommendations
 - Periodicity Worksheet Answer Key User Reviews and Ratings
 - Periodicity Worksheet Answer Key and Bestseller Lists
5. Accessing Periodicity Worksheet Answer Key Free and Paid eBooks
 - Periodicity Worksheet Answer Key Public Domain eBooks
 - Periodicity Worksheet Answer Key eBook Subscription Services
 - Periodicity Worksheet Answer Key Budget-Friendly Options

6. Navigating Periodicity Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodicity Worksheet Answer Key Compatibility with Devices
 - Periodicity Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodicity Worksheet Answer Key
 - Highlighting and Note-Taking Periodicity Worksheet Answer Key
 - Interactive Elements Periodicity Worksheet Answer Key
8. Staying Engaged with Periodicity Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodicity Worksheet Answer Key
9. Balancing eBooks and Physical Books Periodicity Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodicity Worksheet Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodicity Worksheet Answer Key
 - Setting Reading Goals Periodicity Worksheet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodicity Worksheet Answer Key
 - Fact-Checking eBook Content of Periodicity Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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