Name:

 Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.
- Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu K Ni (Br)

- a. Explain why you made these choices: All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.
- Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.
- 4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

5.

- a. O C (Bc) Nc Same Period
 b. Na Rb (Fr) H Same Group
 c. (Pb) C Sn Si Same Group
 d. Au W S (Fr) Nc Zn Challenge
- For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:
 - a. O C Bc No Same Period
 b. Na Rb Fr H Same Group
 c. Pb C Sn Si Same Group
 d. Au W S Fr No Zn Challenge

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Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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