

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

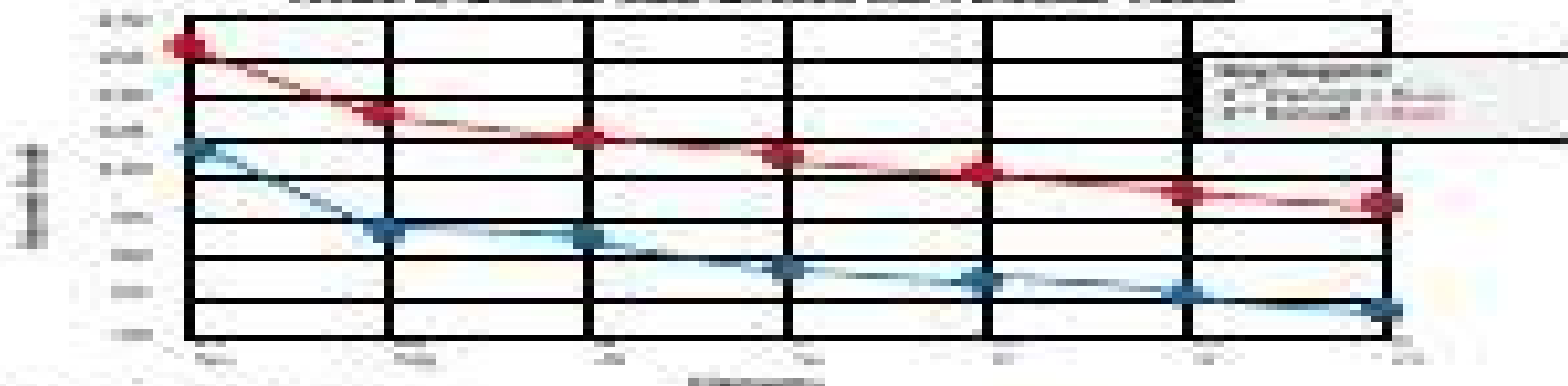
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not, fill in the value graph
3. Label the box as 'not trended' or 'trended'
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	1H	2He	3Li	4Be	5B	6C	7N	8O
2	9F	10Ne	11Na	12Mg	13Al	14Si	15P	16S
3	17Cl	18Ar	19K	20Ca	21Sc	22Ti	23V	24Cr
4	25Mn	26Fe	27Co	28Ni	29Cu	30Zn	31Ga	32Ge
5	33As	34Se	35Br	36Kr	37Rb	38Sr	39Y	40Zr
6	41Nb	42Mo	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd
7	49In	50Sn	51Sb	52Te	53I	54Xe	55Ba	56La

Now label the trends in the following table:

Trend in Atomic Size across the Periodic Table



Electronegativity Questions

1. Following the trend in atomic size across the periodic table, decreases
2. This means that atoms that are more electronegative (more electronegative) are more electronegative.

Electronegativity Questions

1. Electronegativity increases from left to right, i.e. from Na to Cl. This is the electronegativity trend across the periodic table.
2. This means that atoms that are more electronegative (more electronegative) are more electronegative.

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Annelies Wilder-Smith



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The Journal of the Oklahoma State Medical Association ,1991

Ecology Abstracts ,1988

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