

SCPS Chemistry Worksheet – Periodicity

A. Periodic table

1. Which are metals? Circle your answers:



Which metal in the list above has the most metallic character? Explain.

Cesium – as the largest atom, the lowest ionization energy and the most reactivity with nonmetals. This can be determined by its position lowest in the alkali metal group.

2. Write the charge that each of the following atoms will have when it has a complete set of valence electrons forming an ion.

O – 2- Na 1+ F 1- N 3- Ca 2+ Ar - none

3. What is the most common oxidation number for calcium? Explain.

The last page of the powerpoint on *Chemistry: Atoms and Ions* provides the definition of oxidation number as the charge on the ion. You were also asked to define this term in your element brochure. Pages 222 and 980 of your text book defines oxidation number as the positive or negative charge of a monatomic (one atom) ion. These numbers may be confirmed on many of the websites you have used.

Calcium will have a 2+ oxidation number as it tends to easily lose its 2 valence electrons, the 4s² electrons.

4. Name two more elements with that oxidation number and explain your choice.

Other alkaline earth metals (group 2) will also have 2+ oxidation numbers since they all have 2 s valence electrons that they easily lose. This would include Be, Mg, Sr, Ba and Ra.

5. What element in period 3 is a metalloid?

Silicon – Si is the only metalloid in period three.

6. When element with atomic number 118 is discovered, what family will it be in?

Element 118 should be a noble gas, as 118 electrons would arrange to fill 7p sublevel.

7. Make an argument for placing hydrogen in the halogen family rather than the alkali metals.

Since Hydrogen – H has 1 valence electron it is usually placed in the alkali metal group, as that electron is found in the s sublevel. However, the 1s sublevel has room for only 2 electrons. This creates a situation where the sublevel needs only 1 more electron to be complete. Members of the halogen family also only need one electron to complete their sublevel (p). Therefore, hydrogen would fit in this group as well – if the group was defined as needing only one more electron to complete a sublevel rather than the s AND p sublevels.

8. The **alkali metals** have a single electron in the highest energy level.

9. The **alkaline earth metals** achieve the electron configurations of noble gases by losing 2 e-

10. The **transition metals** vary in the number of electrons in the highest energy level.

11. The **halogens** achieve the electron configuration of noble gases by gaining one electron.

12. The **noble gases** have full s and p orbitals in the highest occupied energy levels.

13. The **noble gases** are stable and un-reactive (**inert**).

14. The **halogens** are highly reactive and readily form salts with metals.

15. The **alkaline earth metals** are metals that are more reactive than the transition elements but less reactive than the alkali metals.

16. Predict the oxidation number based on the electron configuration shown.

1s² 2s² 2p⁶ 3s² (2+)

1s² 2s² 2p⁶ 0 (none)

1s² 2s² 2p⁶ 3s¹ (1+)

1s² 2s² 2p⁶ (1-)

1s² 2s² 2p¹ (3+)

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J Rink



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