

Study Guide: Periodic Table and Periodic Trends (Honors)

1. Differentiate between Mendeleev's periodic table and Moseley's periodic table.

Mendeleev put the elements in order by atomic mass, but he noticed that in several places, the atomic masses did not always increase sequentially. Moseley discovered that ordering by atomic number was the correct way to describe the order.

2. State the periodic law.

The properties of the elements are periodic functions of their atomic numbers.

3. Vertical columns on the periodic table are called **groups or families** and horizontal rows are called **periods**.

4. Elements with similar chemical properties are placed in the same **group**.

5. Complete the table below.

Electron Config.	Block	Period	Group/Type	Element
[Xe] 6s ² 4f ¹⁴ 5d ¹⁰ 6p ⁶	p	6	18 / Noble Gas	Radon
[Xe] 6s ¹	s	6	1 / Alkali Metal	Cesium
[Ar] 4s ¹ 3d ⁵	d	4	6 / Trans. Metal	Chromium
[He] 2s ²	s	2	2 / Alk. Earth Metal	Beryllium
[Kr] 5s ¹ 4d ¹⁰	d	5	11 / Trans. Metal	Silver
[He] 2s ² 2p ⁵	p	2	17 / Halogen	Fluorine

6. Determine whether the following statements are describing alkali metals (A), alkaline earth metals (AE), transition metals (TM), halogens (H), noble gases (NG):

- a) most reactive non-metals **H**
- b) s²p⁶ valence configuration **NG**
- c) second most reactive metals **AE**
- d) s¹ valence configuration **A**
- e) exceptions found to e- configurations **TM**
- f) soft metals **A**
- g) d orbitals being filled **TM**

7. Would you expect Li and S to have different chemical properties? Why? **Yes; Li is an alkali metal with 1 valence e-, and S is a nonmetal with 6 valence e-.**

8. Explain why a magnesium atom is smaller than: a sodium atom; a calcium atom. **Mg is smaller than Na because the greater nuclear charge in Mg pulls the e- cloud in tighter (periodic trend). Mg is smaller than Ca because it has fewer energy levels (group trend).**

9. Predict the size of the astatine atom compared to the sulfur atom. **At is larger than S, because At is 3 periods down from S, but only 1 group over. The trends compete, but the sizeable increase in energy levels overcomes the slight difference in nuclear charge.**

10. Would you expect a Cl⁻ ion to be larger or smaller than an Mg²⁺ ion. Explain. **A Cl⁻ ion is larger than a Mg²⁺ ion, because the chloride ion has 3 energy levels, matching**

Periodic Trends Study Guide

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