

# CHAPTER 5: PERIODIC

## The Periodic Law

### SECOND PART

Student's Assistant Answered the following questions for the second part of the lab.

1. Consider the three elements below with  $10^3$  protons each. The numbers in parentheses are atomic numbers.

- \_\_\_\_\_ (a. The first is the noble gas, Xenon?)
- \_\_\_\_\_ (b. The first is the noble gas, Xenon?)
- \_\_\_\_\_ (c. The element is periodic in a noble gas periodic table determined by the atomic number as by the atomic mass?)

2. Consider the element with the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^4$ .

- \_\_\_\_\_ (a. The first period above the noble gas, Xenon?)
- \_\_\_\_\_ (b. It is in a halogen, what is the next noble gas?)
- \_\_\_\_\_ (c. The group number will equal  $10 + 2 + 4 = 16$ . That is, 16th?

3. The noble gases are chemically inert, e.g.,  $He$ ,  $Ne$ ,  $Ar$ .

- \_\_\_\_\_ (a. The noble gases are chemically inert, e.g.,  $He$ ,  $Ne$ ,  $Ar$ .)
- \_\_\_\_\_ (b. The noble gases are other elements, chemical inertness, e.g.,  $He$ ,  $Ne$ ,  $Ar$ .)
- \_\_\_\_\_ (c. Noble gas atom is chemically inert, e.g.,  $He$ ,  $Ne$ ,  $Ar$ .)
- \_\_\_\_\_ (d. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (e. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)

4. Noble gas configuration periodic table, the first row of the periodic table is the noble gas configuration.

- \_\_\_\_\_ (a. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (b. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (c. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (d. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (e. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (f. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (g. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (h. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (i. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (j. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (k. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (l. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (m. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (n. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (o. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (p. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (q. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (r. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (s. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (t. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (u. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (v. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (w. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (x. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (y. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)
- \_\_\_\_\_ (z. The noble gas configuration is  $1s^2 2s^2 2p^6 3s^2 3p^4$ .)

# Periodic Law Mixed Review Answer Key

**Guillaume Favre**

A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the author's name.

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