



Percent Yield

45 grams of nitrogen monoxide (NO) react with an excess of sodium hydroxide (NaOH) following the balanced equation above. 6 grams of water were obtained from the reaction

What is the % yield?

Nitrogen monoxide (NO) gmm:

$$(1)(14.0) + (1)(16.0) = 30.0 \text{ g/mol}$$

$$\text{Actual Yield} = 6 \text{ g H}_2\text{O}$$

Water (H₂O) gmm:

$$(2)(1.0) + (1)(16.0) = 18.0 \text{ g/mol}$$

$$\text{Theoretical Yield} = 6.75 \text{ g H}_2\text{O}$$

$$\frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100 = 88.9\%$$

$$\frac{45 \text{ g NO} \times \frac{1 \text{ mol NO}}{30 \text{ g NO}} \times \frac{4 \text{ mol NO}}{4 \text{ mol NO}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \times 18 \text{ g H}_2\text{O}}{1} = 6.75 \text{ g H}_2\text{O}$$

Percent Yield Problems Answers

K Morrison



Percent Yield Problems Answers:

Unveiling the Energy of Verbal Art: An Psychological Sojourn through **Percent Yield Problems Answers**

In a world inundated with monitors and the cacophony of quick communication, the profound power and mental resonance of verbal beauty frequently fade into obscurity, eclipsed by the constant onslaught of noise and distractions. Yet, located within the lyrical pages of **Percent Yield Problems Answers**, a charming perform of fictional brilliance that pulses with natural emotions, lies an unforgettable trip waiting to be embarked upon. Published with a virtuoso wordsmith, this magical opus manuals viewers on an emotional odyssey, lightly exposing the latent possible and profound impact stuck within the elaborate internet of language. Within the heart-wrenching expanse of the evocative analysis, we shall embark upon an introspective exploration of the book is main themes, dissect its interesting writing type, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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