

# Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before  $\underline{2} \text{ mol}$   $\underline{0} \text{ mol}$   $\underline{0} \text{ mol}$

Change  $\underline{-2 \text{ mol}}$   $\underline{+2 \text{ mol}}$

After  $\underline{0 \text{ mol}}$   $\underline{2 \text{ mol}}$

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

- Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before  $\underline{9.94 \text{ mol}}$   $\underline{x \text{ mol}}$   $\underline{0 \text{ mol}}$

Change  $\underline{-9.94 \text{ mol}}$   $\underline{+9.94 \text{ mol}}$

After  $\underline{0 \text{ mol}}$   $\underline{9.94 \text{ mol}}$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

- You eat 180.0 g of glucose (90 M&Ms). If glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ , reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to

# Percent Yield Stoichiometry Answers With Work

**KJ Lindholm-Leary**



## **Percent Yield Stoichiometry Answers With Work:**

## **Percent Yield Stoichiometry Answers With Work** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the power of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such may be the essence of the book **Percent Yield Stoichiometry Answers With Work**, a literary masterpiece that delves deep in to the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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