Chapter 9

9.1 Compounds and Chemical Change

- Atom- the smallest unit of an element that can exist alone or in combination with other elements.
- Compounds- formed when atoms are held together by an attractive force called a chemical bond
- Molecule- The smallest particle of a compound or a gaseous element, that can exist and still remain the characteristic chemical properties of a substance
- Chemical reaction

 a change in matter where different chemical substances are created by breaking and/or forming chemical bonds.
 - Examples: Photosynthesis
- Energy is absorbed -> produce new chemical substances -> more internal potential energy
- Energy is released when new chemical substances are produced -> less internal potential energy
- Chemical Equation: a way of describing what happens during a chemical reaction
 - Example: Photosynthesis equation-

Energy/Sunlight + carbon dioxide molecules + water molecules -> plant material molecules + oxygen molecules

9.2 Valence Electrons and Ions

- the number of electrons in the outermost orbital that usually determines the chemical properties of an atom
- <u>Valence electrons</u> the number of electrons in the outermost orbital that usually determines the properties of an atom
- Octet Rule after the element helium, atoms attempt to acquire an outer orbital with eight electrons through chemical reactions

9.3 Chemical Bonds

- chemical bond-an attractive force that holds atoms together in a compound.
- Three Classes of Chemical Bonds-
 - Ionic Bonds
 - Covalent Bonds
 - Metallic Bonds
- Fonic Bonds
 - type of linkage formed from the electrostatic attraction between oppositely charged ions in a chemical compound
- Covalent bonds: result when atoms achieve the noble gas electron structure by sharing
 electrons. They generally formed between the nonmetallic elements on the right side of
 the periodic table
- Metallic bonds: formed in solid metals such as iron, copper, and the other metallic elements that make up about 80 percent of all the elements.
- . Heat formation- energy that is released during an ionic bond formation
- Boemula- describes what elements are in the compound and in what proportions
- Simple bond covalent bond
- Double bond- formed by two pairs of electrons that are shared by two atoms

Physical Science Chapter Vocabulary

Kaye Hagler

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Table of Contents Physical Science Chapter Vocabulary

- 1. Understanding the eBook Physical Science Chapter Vocabulary
 - The Rise of Digital Reading Physical Science Chapter Vocabulary
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Chapter Vocabulary
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter Vocabulary
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Chapter Vocabulary
 - Personalized Recommendations
 - Physical Science Chapter Vocabulary User Reviews and Ratings
 - Physical Science Chapter Vocabulary and Bestseller Lists
- 5. Accessing Physical Science Chapter Vocabulary Free and Paid eBooks
 - Physical Science Chapter Vocabulary Public Domain eBooks
 - Physical Science Chapter Vocabulary eBook Subscription Services
 - Physical Science Chapter Vocabulary Budget-Friendly Options

- 6. Navigating Physical Science Chapter Vocabulary eBook Formats
 - o ePub, PDF, MOBI, and More
 - Physical Science Chapter Vocabulary Compatibility with Devices
 - Physical Science Chapter Vocabulary Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter Vocabulary
 - Highlighting and Note-Taking Physical Science Chapter Vocabulary
 - Interactive Elements Physical Science Chapter Vocabulary
- 8. Staying Engaged with Physical Science Chapter Vocabulary
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter Vocabulary
- 9. Balancing eBooks and Physical Books Physical Science Chapter Vocabulary
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter Vocabulary
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chapter Vocabulary
 - Setting Reading Goals Physical Science Chapter Vocabulary
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chapter Vocabulary
 - Fact-Checking eBook Content of Physical Science Chapter Vocabulary
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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