

Periodic Trends Worksheet

Directions: Use the table to complete the following questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

Na, K, S, O, Cl

3. Which atom has the highest first ionization energy than sodium?

Chlorine (because it has the most electronegativity and is closest to sodium)

4. Which two elements in the same family generally have similar properties?

Sulfur and chlorine (because they are adjacent & share similar outer electrons)

5. Arrange **halogens** the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (decreasing **left to right**) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms of the periodic table? Write names, this period?

Period 4, because it contains elements ranging from potassium to the noble gas krypton.

7. Which trend in ionization energy atoms across a period and the periodic table? Which element shows this?

Electron configuration (increasing electronegativity, because it is closest to sodium)

8. Write the atoms include each their respective atomic radius.

Li	152
Be	112
B	85
C	77
N	75
O	73
F	72
Ne	71

Na	186
Mg	145
Al	118
Si	111
P	106
S	104
Cl	99
Ar	94

9. Write the atoms include each their respective electronegativity.

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Table of Contents Physical Science If8767 Pg 63 Key

1. Understanding the eBook Physical Science If8767 Pg 63 Key
 - The Rise of Digital Reading Physical Science If8767 Pg 63 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Pg 63 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science If8767 Pg 63 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Pg 63 Key
 - Personalized Recommendations
 - Physical Science If8767 Pg 63 Key User Reviews and Ratings
 - Physical Science If8767 Pg 63 Key and Bestseller Lists
5. Accessing Physical Science If8767 Pg 63 Key Free and Paid eBooks
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 - ePub, PDF, MOBI, and More
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 - Physical Science If8767 Pg 63 Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Physical Science If8767 Pg 63 Key
 - Interactive Elements Physical Science If8767 Pg 63 Key
- 8. Staying Engaged with Physical Science If8767 Pg 63 Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science If8767 Pg 63 Key
- 9. Balancing eBooks and Physical Books Physical Science If8767 Pg 63 Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science If8767 Pg 63 Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science If8767 Pg 63 Key
 - Setting Reading Goals Physical Science If8767 Pg 63 Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science If8767 Pg 63 Key
 - Fact-Checking eBook Content of Physical Science If8767 Pg 63 Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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