



Class 10 Physical Science

Chemistry Formal Experiment: *Using flame colours to identify unknown metal ions*

August 2024

60 Marks

1 hour

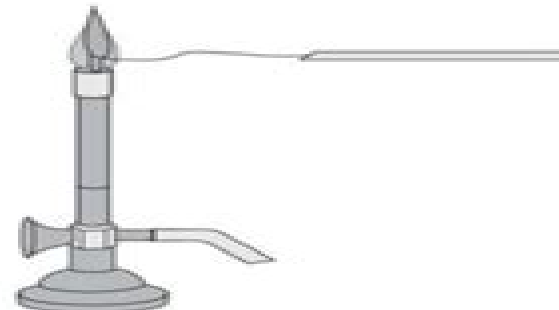
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Introduction:

The flame test is used to visually determine the identity of an unknown metal or metalloid, as an element or in a compound, based on the characteristic colour the salt turns the flame of a burner. When an electron in an excited state (higher energy) returns to its ground state (lower energy), an atomic spectrum is produced. A specific quantity of energy is associated with each electron, and when an electron returns to a lower energy level, a specific quantity of energy is released. The energy is released as electromagnetic radiation. Certain wavelengths will be visible to the naked eye; other wavelengths can only be detected with instruments. In the activity below, we will investigate the colours emitted by certain metal atoms. Different colours indicate different wavelengths of visible radiation.

Materials:

- ☐ NaCl(aq) , KCl(aq) , $\text{CaCl}_2 \cdot 2\text{H}_2\text{O(aq)}$, $\text{CuCl}_2 \cdot 2\text{H}_2\text{O (aq)}$,
- ☐ copper powder, magnesium powder,
zinc powder, iron powder
- ☐ four unknown salt solutions
- ☐ Bunsen burner
- ☐ Nichrome wire



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