

ATOMIC STRUCTURE AND THE PERIODIC TABLE CHAPTER 4 WORKSHEET

PART A

Given the following isotopes, determine the atomic number, the mass number, the number of protons, electrons and neutrons.

Isotope Symbol	Atomic Number	Mass Number	Protons	Electrons	Neutrons	Isotope Name
$^{131}_{53}\text{I}$	53	131	53	53	78	iodine - 131
$^{35}_{16}\text{S}$	16	35	16	16	19	sulphur - 35
^4_2He	2	4	2	2	2	helium - 4
$^{27}_{13}\text{Al}$	13	27	13	13	14	aluminium - 27
$^{81}_{36}\text{Kr}$	36	81	36	36	45	krypton - 81
$^{81}_{37}\text{Rb}$	37	81	37	37	44	rubidium - 81

PART B

Complete the following chart by writing the symbol for the isotope of the following elements. In addition, give the number of protons, electrons, mass number and atomic number and complete the element name.

Element Name	Neutrons	Protons	Electrons	Mass Number	Atomic Number	Isotope Symbol
Uranium- 237	145	92	92	237	92	$^{237}_{92}\text{U}$
Chlorine- 45	28	17	17	45	17	$^{45}_{17}\text{Cl}$
Oxygen- 17	9	8	8	17	8	$^{17}_8\text{O}$
Boron- 11	6	5	5	11	5	$^{11}_5\text{B}$
Beryllium- 9	5	4	4	9	4	^9_4Be
Hydrogen- 2	1	1	1	2	1	^2_1H
Carbon- 14	8	6	6	14	6	$^{14}_6\text{C}$

Physical Science Chapter 4 Answers

R. M. Harbeck



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Table of Contents Physical Science Chapter 4 Answers

1. Understanding the eBook Physical Science Chapter 4 Answers
 - The Rise of Digital Reading Physical Science Chapter 4 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 4 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 4 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 4 Answers
 - Personalized Recommendations

- Physical Science Chapter 4 Answers User Reviews and Ratings
- Physical Science Chapter 4 Answers and Bestseller Lists
- 5. Accessing Physical Science Chapter 4 Answers Free and Paid eBooks
 - Physical Science Chapter 4 Answers Public Domain eBooks
 - Physical Science Chapter 4 Answers eBook Subscription Services
 - Physical Science Chapter 4 Answers Budget-Friendly Options
- 6. Navigating Physical Science Chapter 4 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 4 Answers Compatibility with Devices
 - Physical Science Chapter 4 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 4 Answers
 - Highlighting and Note-Taking Physical Science Chapter 4 Answers
 - Interactive Elements Physical Science Chapter 4 Answers
- 8. Staying Engaged with Physical Science Chapter 4 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 4 Answers
- 9. Balancing eBooks and Physical Books Physical Science Chapter 4 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 4 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chapter 4 Answers
 - Setting Reading Goals Physical Science Chapter 4 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chapter 4 Answers
 - Fact-Checking eBook Content of Physical Science Chapter 4 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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