

Periods and Groups: Valence Electrons: Memory

Periods and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group										18A	
	1A	2A	3A	4A	5A	6A	7A	8A	9A	10A	11A	12A	13A	14A
1	H	He											Ne	
2	Li	Be	B	C	N	O	F	Ne					Ar	
3	Na	Mg	Al	Si	P	S	Cl	Ar					Kr	
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge

Transition metals in between

Thinking like the game "Memory"? ...remember that an element's in period 2 and group 1.

What element is in Period 2, and Group 16A? (Oxygen) ...what is Group 16A in Column 16A? (The chalcogen group)

What element is in Period 3 and Group 1?

What element is in Group 2A, and Period 3?

What group would find a halogen?

Group: _____ Period: _____

What group halogens?

Group: _____ Period: _____

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each period contains the same number of valence electrons. Valence electrons are the electrons that get involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Lithium 3



3 valence electrons

Hydrogen and Helium are built in Group 1A-2A. Helium is 2 valence electrons.

Group 1A, Helium 18A and the 18 valence electrons. How many valence electrons will they have?

How many electrons?

Group 18A - How many more shells that have a full shell of electrons. Elements are in Group 18 and Helium are built in 18 shells - 18. Helium has 2 valence electrons. "He 2 is, 18A" Group 18A has a full shell of valence electrons. Helium 18A has 2 valence electrons. Helium 18A has 2 valence electrons.

Periods And Groups Valence Electrons Masses Answers

Jin-Ying Zhang



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