

# Graphing Periodic Trends Lab

Name \_\_\_\_\_ Hr \_\_\_\_\_

**PURPOSE:** To determine how certain properties are periodic when the elements are arranged in periods or groups.

**DISCUSSION:** There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

## REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

Ionization energy - The energy required to remove an electron from an atom.

## DATA

**Properties of the First 18 Elements**

Element	Atomic Number	Atomic Radius (Angstroms)	Ionization Energy
H	1	0.79	314
He	2	0.49	567
Li	3	2.05	124
Be	4	1.4	215
B	5	1.17	191
C	6	0.91	260
N	7	0.75	335
O	8	0.65	314
F	9	0.57	402
Ne	10	?	?
Na	11	2.23	119
Mg	12	1.72	176
Al	13	1.82	138
Si	14	1.46	168
P	15	1.23	242
S	16	1.09	239
Cl	17	0.97	299
Ar	18	?	?

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with question marks.

1. Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)
2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)

# Periodic Trends Lab Answers

**R Sanford**



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