

5.1 Atomic structure and the Periodic Table

You will learn:

- To understand how the position of an element in the Periodic Table can be used to predict its atomic structure and chemical and physical properties.
- To understand that models and analogies can change according to scientific evidence
- To describe the uses, strengths and limitations of models and analogies
- To represent scientific ideas using recognised symbols or formulae

1.

Draw a line to match the word with its correct meaning.

The first line has been drawn for you.

| Word | Meaning |
|-----------|-----------------------------------------------------------|
| Electrons | Particles with no charge found in an atom |
| Neutrons | The central part of the atom |
| Protons | The paths or orbits that electrons move along in the atom |
| Nucleus | Very small negatively charged particles in an atom |
| Shells | Particles with a positive charge found in an atom |

2.

Complete the sentences.

Use the words in the box.

| | | | | |
|--------|--------|----------|-----------|---------|
| atomic | mass | protons | electrons | nucleus |
| | shells | neutrons | atom | |

- a An _____ number is the protons an _____ of an element has.
- b A _____ number is the total number of _____ and _____ found in the nucleus of an atom.

Periodic Table Workbook Answers Chemistry

M Walker



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