

Quiz & Worksheet - Atomic and Ionic Radii on the Periodic Table

1. As you move from left to right on the periodic table, why does the atomic radius decrease?

- ☐ As more protons are added, the more tightly the surrounding electrons are drawn inwards.
- ☐ As more neutrons are added, the more tightly the surrounding electrons are drawn inwards.
- ☐ As more electrons are added, the more tightly they draw themselves towards the nucleus.
- ☐ As the mass of the nucleus rises, the more tightly the surrounding electrons are drawn inwards.

2. As you move from top to bottom on the periodic table, what happens to the atomic radius?

- ☐ It decreases.
- ☐ It decreases, then increases.
- ☐ It remains constant.
- ☐ It increases, then decreases.
- ☐ It increases.

3. Which of these elements has the largest atomic radius: argon, phosphorus, aluminum, sodium or magnesium?

A standard periodic table of elements. The elements are arranged in rows and columns, with their atomic numbers and chemical symbols displayed in each cell. The table includes all elements from Hydrogen (H) to Oganesson (Og).

- ☐ Argon (Ar)
- ☐ Phosphorus (P)
- ☐ Sodium (Na)
- ☐ Aluminum (Al)
- ☐ Magnesium (Mg)

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George Graybill



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