



Percent Yield

45 grams of nitrogen monoxide (NO) react with an excess of sodium hydroxide (NaOH) following the balanced equation above. 6 grams of water were obtained from the reaction

What is the % yield?

Nitrogen monoxide (NO) gmm:

$$(1)(14.0) + (1)(16.0) = 30.0 \text{ g/mol}$$

Water (H₂O) gmm:

$$(2)(1.0) + (1)(16.0) = 18.0 \text{ g/mol}$$

$$\text{Actual Yield} = 6 \text{ g H}_2\text{O}$$

$$\text{Theoretical Yield} = 6.75 \text{ g H}_2\text{O}$$

$$\frac{6}{6.75} \times 100 = 88.9\%$$

45 g NO	1 mol NO	1 mol H₂O	18 g H ₂ O
<hr/>			
30 g NO	4 mol NO	1 mol H₂O	= 6.75 g H ₂ O



Percent Yield Problems Answers

Joacim Rocklöv



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