

If 16.0 g of  $\text{CaCO}_3$  are decomposed and 7.54 g of  $\text{CaO}$  are obtained, what is the percent yield of  $\text{CaO}$  obtained?  $\leftarrow$  A.Y



$$\text{Percent Yield} = \frac{\text{Actual yield}}{\text{theoretical yield}} \times 100$$

$$16\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.1\text{g CaCO}_3} \times \frac{1\text{mol CaO}}{1\text{mol}}$$

# Percent Actual And Theoretical Yield Key

**SJ Ball**



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