

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Moles of product based on the limiting reagent. Then convert moles of product to grams. Percent yield = (actual yield / theoretical yield) x 100.

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
 Calculate: Moles of product based on the limiting reagent. Then convert moles of product to grams.



What mass of water can be produced from 2.00 g of H_2 ?

2.00 g H_2	18.0 g H_2O
2.00 g H_2	18.0 g H_2O
2.00 g H_2	18.0 g H_2O
2.00 g H_2	18.0 g H_2O

- 2.00 g H_2 is the limiting reagent. 18.0 g H_2O is the theoretical yield. 18.0 g H_2O is the actual yield. 100% yield.



- 2.00 g H_2 is the limiting reagent. 18.0 g H_2O is the theoretical yield. 18.0 g H_2O is the actual yield. 100% yield.

- 18.0 g H_2O is the limiting reagent. 18.0 g H_2O is the theoretical yield. 18.0 g H_2O is the actual yield. 100% yield.



- 18.0 g H_2O is the limiting reagent. 18.0 g H_2O is the theoretical yield. 18.0 g H_2O is the actual yield. 100% yield.



Percent Yield And Limiting Reactants Quiz

SA Adler



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