

PHYSICAL SCIENCES – GRADE 10

CHAPTER – 1 Chemical Reactions and Equations

- q **Chemical Reaction** : – Whenever a chemical change occurs we can say that a chemical reaction has taken place
eg – Food gets digested in our body
 – Rusting of iron.
- q **Chemical Equation** :– A chemical reaction can be expressed symbolically by using chemical equation
eg magnesium is burnt into air to form magnesium oxide can be represented as
$$\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$$

– We can observe or recognise a chemical reaction by observing change in state, colour, by evolution of gas or by change in temperature.
- q Physical state of the reactant and products are mentioned to make chemical reaction more informative. eg we use (g) for gas, (l) for liquid, (s) for solid and (aq) for aqueous.
- q **Balancing Equation** :– We balance the chemical equation so that no. of atoms of each element involved in the reaction remain same at the reactant and product side.
eg $\text{Fe} + \text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + \text{H}_2$ can be written as
$$3\text{Fe(s)} + 4\text{H}_2\text{O(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)} + 4\text{H}_2\text{(g)}$$
- q **Combination Reaction** :– The reaction in which two or more substances combine to form a new single substance

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Albert A Gayle



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Table of Contents Physical Science P2 Final Exam Grade 10 2014

1. Understanding the eBook Physical Science P2 Final Exam Grade 10 2014
 - The Rise of Digital Reading Physical Science P2 Final Exam Grade 10 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science P2 Final Exam Grade 10 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science P2 Final Exam Grade 10 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science P2 Final Exam Grade 10 2014
 - Personalized Recommendations
 - Physical Science P2 Final Exam Grade 10 2014 User Reviews and Ratings
 - Physical Science P2 Final Exam Grade 10 2014 and Bestseller Lists
5. Accessing Physical Science P2 Final Exam Grade 10 2014 Free and Paid eBooks
 - Physical Science P2 Final Exam Grade 10 2014 Public Domain eBooks
 - Physical Science P2 Final Exam Grade 10 2014 eBook Subscription Services
 - Physical Science P2 Final Exam Grade 10 2014 Budget-Friendly Options
6. Navigating Physical Science P2 Final Exam Grade 10 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science P2 Final Exam Grade 10 2014 Compatibility with Devices
 - Physical Science P2 Final Exam Grade 10 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science P2 Final Exam Grade 10 2014
 - Highlighting and Note-Taking Physical Science P2 Final Exam Grade 10 2014
 - Interactive Elements Physical Science P2 Final Exam Grade 10 2014

8. Staying Engaged with Physical Science P2 Final Exam Grade 10 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science P2 Final Exam Grade 10 2014
9. Balancing eBooks and Physical Books Physical Science P2 Final Exam Grade 10 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science P2 Final Exam Grade 10 2014
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science P2 Final Exam Grade 10 2014
 - Setting Reading Goals Physical Science P2 Final Exam Grade 10 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science P2 Final Exam Grade 10 2014
 - Fact-Checking eBook Content of Physical Science P2 Final Exam Grade 10 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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