(2)

(1)

(3) [13]

QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The letters A, B and C represent the compounds CH₄, C₄H₆ and C₃H₈ respectively.



- 2.1.1 Define the term boiling point.
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph.
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (A, B or C) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph.
 (2)
- 2.2 Consider the two molecules in the table below.

NAME OFSUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT
Ammonia	NH ₃	17	- 33
Phosphine	PH ₃	×	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES.

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- 1. Understanding the eBook Physical Science Grade Question Papers
 - The Rise of Digital Reading Physical Science Grade Question Papers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Grade Question Papers
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - $\circ\,$ Features to Look for in an Physical Science Grade Question Papers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Grade Question Papers
 - Personalized Recommendations
 - Physical Science Grade Question Papers User Reviews and Ratings
 - Physical Science Grade Question Papers and Bestseller Lists
- 5. Accessing Physical Science Grade Question Papers Free and Paid eBooks
 - Physical Science Grade Question Papers Public Domain eBooks
 - Physical Science Grade Question Papers eBook Subscription Services
 - Physical Science Grade Question Papers Budget-Friendly Options

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 - Physical Science Grade Question Papers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Physical Science Grade Question Papers
 - Highlighting and Note-Taking Physical Science Grade Question Papers
 - Interactive Elements Physical Science Grade Question Papers
- 8. Staying Engaged with Physical Science Grade Question Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade Question Papers
- 9. Balancing eBooks and Physical Books Physical Science Grade Question Papers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Physical Science Grade Question Papers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Grade Question Papers
 - Setting Reading Goals Physical Science Grade Question Papers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Grade Question Papers
 - Fact-Checking eBook Content of Physical Science Grade Question Papers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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mephistopheles will not answer this we see faustus becoming wracked with regret he finally cries the name of christ and at this lucifer and beelzebub join them faustus presumes they have come to fetch his soul they tell him that they have come to faustus dr faustus dr faustus dr faustus gce drama 2016 dr faustus

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