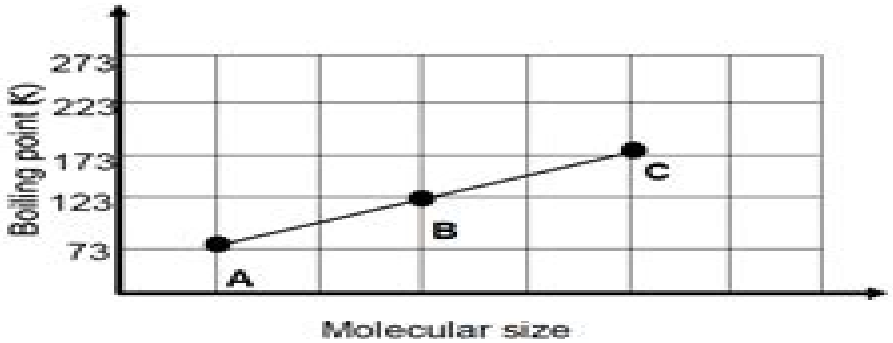
QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.
The **letters A, B** and **C** represent the compounds CH_4, C_4H_8 and C_3H_8 respectively.



2.1.1 Define the term *boiling point*. (2)

2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)

2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

2.1.4 Which ONE of the compounds (**A, B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)

2.2 Consider the two molecules in the table below.

NAME OFSUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT (°C)
Ammonia	NH_3	17	- 33
Phosphine	PH_3	x	– 87,4

2.2.1 Determine the molecular mass (x) of phosphine (1)

2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

[13]

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