

Periodicity - Electronegativity

Name: Answer Key

1. Fill in the blanks below, and then draw arrows on to the periodic table indicating the direction of the electronegativity trend.

Generally, electronegativity

increases

across periods and

increases

up groups.

↑
increases

→
increases

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100
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2. List the elements below in order of increasing electronegativity. (Smallest to largest).

Ba, Rb, Ca, Se, Mg

Rb, Ba, Ca, Mg, Se

Zn, Ni, Cu, Sc, Ti

Sc, Ti, Ni, Cu, Zn

F, B, O, Rb, Fr

Fr, Rb, B, O, F

Nb, Sn, Rb, Fr, I

Fr, Rb, Nb, Sn, I

F, Li, Ba, K, Na

K, Na, Li, Ba, F

P, Se, Ar, Mg, Ca

Ca, Mg, Se, P, Ar

3. Are the elements below in order of decreasing electronegativity? If yes, write "yes". If not, write them in the correct order.

Ca, K, Se, Br

Br, Se, Ca, K

I, Br, Cl, F

F, Cl, Br, I

O, B, Li, Na

Yes

Y, Se, Rb, Fr

Yes

At, Sc, Se, Ba

Yes

As, Pb, Te, Si

Si, Te, Pb, As

4. Briefly explain why electronegativity changes up groups.

Generally, electronegativity increases as you move up in a group. This is because the atomic number decreases, and as a result, there is less distance between the nucleus and valence electrons. In other words, there is a smaller atomic radius.

5. Briefly explain why electronegativity changes across periods.

As you move from left to right, across a period, the nuclear charge of elements increases faster than electron shielding. This results in an increased attraction to electrons.

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