

Periodic Trends: Straw Lab

Objective: To create a visual representation of the following periodic trends: atomic size, ionization energy and electronegativity.

Atomic Radius: the size of an atom measured in either nanometers (nm) or angstroms (Å).

Ionization Energy: the energy needed to remove an electron from an atom measured in kilojoules per mole (kJ/mol) of an atom.

Electronegativity: the relative ability of an atom to attract electrons towards itself.

Materials: For this lab you will need a sheet with 3 small periodic tables, straws, scissors, a metric ruler, your notebook and access to a hot glue gun.

Procedure:

- 1) Draw and complete the data table for each trend in your lab notebook. Each person should have their own data tables in their own notebooks!
- 2) For each trend, calculate the scale factor according to the instructions. Using the scale factor, measure the straw and cut it to the appropriate size.
- 3) You will then use the glue gun to glue the straw to the periodic table in the box that represents the element's location on the periodic table.
- 4) Answer the questions for each trend in your lab notebook. Everyone should have their own answers in their own notebook!

Atomic Radius:

The radius of the following atoms has been estimated to have the following values:

Name	Symbol	Atomic Number	Size of Radius (nm)	Calculated Straw Length (cm)
Fluorine			0.064	
Chlorine			0.0999	
Bromine			0.114	
Iodine			0.1133	
Sodium			0.186	
Magnesium			0.160	
Aluminum			0.143	
Silicon			0.117	

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Michael Hoelscher



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