

CHAPTER 5: PERIODICITY

The Periodic Law

SECOND PART

Student's Assistant Answered the following questions for the second part of the lab.

1. If you consider the groups of elements with $ns^2 np^2$ configuration, and the elements with $ns^2 np^3$ configuration, answer the following questions:

- _____ a. The first ionization potential?
- _____ b. The first ionization potential?
- _____ c. Is the element in period 4 or 5 having periodic table characteristics? Is the element in period 4 or 5 having periodic table characteristics?

2. Consider the elements with the following configurations: $ns^2 np^2$, $ns^2 np^3$, $ns^2 np^4$.

- _____ a. The first ionization potential?
- _____ b. The first ionization potential?
- _____ c. The group number will equal $(ns^2 np^2) + (ns^2 np^3) + (ns^2 np^4)$?

- _____ d. The first ionization potential?
- _____ e. The first ionization potential?

- _____ f. The first ionization potential?

- _____ g. The first ionization potential?
- _____ h. The first ionization potential?

3. Write the electron configuration for the first four elements of Period 4. Is the element with the following configuration a particular element?

- _____ a. The first ionization potential?
- _____ b. The first ionization potential?
- _____ c. The first ionization potential?
- _____ d. The first ionization potential?
- _____ e. The first ionization potential?
- _____ f. The first ionization potential?
- _____ g. The first ionization potential?
- _____ h. The first ionization potential?

Periodic Law Mixed Review Answer Key

Camilla Rothe

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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